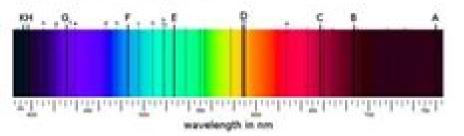
Spectral Analysis



Types of spectra:

Type of spectrum	Photographic example		
Continuous (or continuum)			
Absorption (dark line)			
Emission (bright line)			

Here is a spectrum with many of the known spectral lines.



Designation	Diement	Wavelength (nm)	Designation	Element	Warningth (nm)
,	0.	998,765	34 3	Te.	495,763
Z	O ₁	822,696	P	HB	486,134
A	0,	759.379	34	Fe.	466.814
8	01	686,719		Fe	438.335
C X	He	656.281	0	Ha	434.047
	0:	627,661	G	Per .	430,790
D ₀	No.	589.592	G	-	430,734
D ₁	Na	588.995	- A	154	430.175
Do or d	He.	587,5688	H	Cu'	396.547
• 1	He	546.073	K	Call	299,368
L	74	527,039	L	To .	382.044
be .	Ma	518,342	N	To:	358.121
600 E	Mg	517.276	P	AT7	336,112
lhe .	To .	556.890	T	Fe	302.108
Maria de la companya	Fee	556.181	(F)	100	299,444
h.	Mg	556,755			

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						_
	c. 523 nm		d. 54	60 nm		-
E. W/	net colour is t	he emission o	if hydrog	en-alpha, H	9.7	
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(no	te Some eler	sents have no	one than	one absorpt	ion line)	
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Analysis Of Spectral Lines Answers

D. W. Golightly, Frederick Otto Simon

Analysis Of Spectral Lines Answers:

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Stedman,1872 <u>Chemical News and Journal of Industrial Science</u>,1890 **Lil-'Ulūm Al-Riyāḍīyah Wa-al-Ṭabī'īyah**,1946

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