Balancing redox equations

Example

Balance the equation showing the oxidation of Fe²⁺ ions to Fe³⁺ ions by dichromate ions (Cr₂O₇²⁻) in an acidic medium

$$Fe^{2+} + Cr_2O_7^{2-} \rightarrow Fe^{3+} + Cr^{3+}$$

Step 1: Identify oxidising and reducing agents and write half reactions

Fe²⁺: +2
$$Cr_2O_7^{2-}$$
: $2Cr + 7(-2) = -2$
Fe³⁺: +3 $2Cr = 12$
 Cr^{3+} : +3 $Cr = +6$

$$Fe^{2+} \rightarrow Fe^{3+}$$

Iron loses one electron, i.e. it is oxidised. It donates one electron to dichromate, i.e. it is the reducing agent.

$$Cr_2^{6+} \rightarrow Cr^{3+}$$

Chromium gains three electrons, i.e. it is reduced. It gains its electrons from iron, i.e. it is the oxidising agent.

Half equation: Reduction reaction: Cr26+ + 3e- → Cr3+

Analyzing Oxidation Reduction Reactions Answers

Suzanne Nielsen

Analyzing Oxidation Reduction Reactions Answers:

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creating vertical seismic load manually in staad pro - Oct 04 2022

web i will cover different international codes which are available for static seismic analysis in staad a brief discussion on the

Analyzing Oxidation Reduction Reactions Answers

 $newly \ implemented \ ibc \ 2018 \ code \ different \ options$

tr 32 10 1 response spectrum analysis bentley systems - Mar 09 2023

web jul 24 2019 seismic analysis problem in staad while carrying out seismic analysis by equivalent static method we calculate the forces manually at each floor level

tr 31 2 11 is 1893 part 1 2016 codes lateral seismic - Jul 13 2023

web the staad pro seismic load generator follows the procedure of equivalent lateral load analysis explained in ubc ibc and several other codes it is assumed that the lateral