

Name: Ahmed Faizan Date: 12 May Period:

Specific Heat and Heat Capacity Worksheet

DIRECTIONS: Use $q = (m)(C_p)(\Delta T)$ to solve the following problems. Show all work and units.

Ex: How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$\begin{aligned} q &= ? & \Delta t &= 55 - 22 = 33^\circ\text{C} & q &= m s \Delta t \\ m &= 10\text{g} & s &= 0.9 \text{ J/g}^\circ\text{C} & &= 10 \times 0.9 \times 33 \\ & & & & &= \underline{297 \text{ J}} \end{aligned}$$

1. The temperature of 335 g of water changed from 24.5°C to 26.4°C. How much heat did this sample absorb? C for water = 4.18 J/g°C

$$\begin{aligned} \Delta t &= 26.4 - 24.5 = 2 & q &= m s \Delta t \\ s &= 4.18 \text{ J/g}^\circ\text{C} & &= 335 \times 4.18 \times 2 \\ m &= 335\text{g} & &= \underline{2800.6 \text{ J}} \end{aligned}$$

2. How much heat in kilojoules has to be removed from 225g of water to lower its temperature from 25.0°C to 10.0°C?

$$\begin{aligned} \Delta t &= 25 - 10 = 15 & q &= m s \Delta t \\ m &= 225\text{g} & &= 225 \times 4.18 \times 15 \\ s &= 4.18 \text{ J/g}^\circ\text{C} & &= \underline{14107.5 \text{ J}} \\ 1\text{J} &= 1 \times 10^{-3} \text{ KJ} & \therefore &= \underline{14.1075 \text{ KJ}} \end{aligned}$$

3. To bring 1.0kg of water from 25°C to 99°C takes how much heat input?

$$\begin{aligned} m &= 1\text{kg} = 1000\text{g} & q &= 1000 \times 4.18 \times 74 \\ \Delta t &= 99 - 25 = 74 & q &= \underline{309320 \text{ J}} \\ s &= 4.18 \text{ J/g}^\circ\text{C} & & \end{aligned}$$

An insulated cup contains 75.0g of water at 24.00°C. A 26.00g sample of metal at 82.25°C is added. The final temperature of the water and metal is 28.34°C. What is the specific heat of the metal?

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$$\begin{aligned} (q_{\text{water}} &= -q_{\text{metal}}) & (C = m s) &\rightarrow s = \frac{C}{m} \\ q_{\text{water}} &= 75 \times 4.18 \times (28.34 - 24) & q &= m s \Delta t \\ &= 1252 & 1252 &= 26 \times s \times (82.25 - 28.34) \\ & & & 1252 = 26 \times 53.91 \times s \\ & & & s = \underline{0.893 \text{ J/g}^\circ\text{C}} \end{aligned}$$

5. A calorimeter has a heat capacity of 1265 J/°C. A reaction causes the temperature of the calorimeter to change from 22.34°C to 25.12°C. How many joules of heat were released in this process?

$$\begin{aligned} C &= 1265 \text{ J/}^\circ\text{C} \\ \Delta t &= 25.12 - 22.34 = 2.78 \\ q &= ? \\ q &= C \Delta t \\ &= 1265 \times 2.78 \\ &= \underline{3516.7 \text{ J}} \end{aligned}$$

Answers Of Specific Heat

M Planty



Answers Of Specific Heat:

Answers to Questions Set at Examinations Great Britain. Education Department. Department of Science and Art, 1889

Answers to questions set at the science and art examinations Science and art department, 1887 **Matriculation**

model answers: heat and light, magnetism and electricity. Being the papers from June 1888 to Jan. 1891, with

answers by H.E. Schmitz and R.W. Stewart London univ, exam. papers, 1891 **Fluid and Thermal Dynamics Answer**

Bank for Engineers Ethirajan Rathakrishnan, 2023-03-27 This book provides the essence of aerodynamics fluid mechanics experimental methods gas dynamics high enthalpy gas dynamics helicopter aerodynamics heat transfer and thermodynamics describing the underlying principles of these subjects before listing the set of multiple choice questions of each subject which will prove to be useful for engineering students to comfortably face and win in the competitive examinations for engineering studies engineering services civil services doctoral Degree program entrance and so on This book will also be of value for those facing job interviews for academic positions in universities and research organizations or laboratories Numerical

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Dr. J. P. Goel, Er. Meera Goyal, 2020-06-27 Unit 1 Physical World and Measurement 1 Systems of Units and Measurement 2

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Thermodynamics 20 First Law of Thermodynamics 21 Second Law of Thermodynamics Unit 9 Behaviour of Perfect Gases and

Kinetic Theory of Gases 22 Behaviour of Perfect Gas and Kinetic Theory of Gases Unit 10 Oscillations and Waves 23

Oscillations 24 Speed of Mechanical Waves and Progressive Waves 25 Superposition of Waves Interference and Beats 26

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Lessons in Elementary Physics Balfour Stewart,1895 **Scientific American** ,1887

This book delves into Answers Of Specific Heat. Answers Of Specific Heat is a vital topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Answers Of Specific Heat, encompassing both the fundamentals and more intricate discussions.

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- Chapter 2: Essential Elements of Answers Of Specific Heat
- Chapter 3: Answers Of Specific Heat in Everyday Life
- Chapter 4: Answers Of Specific Heat in Specific Contexts
- Chapter 5: Conclusion

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6. In chapter 5, the author will draw a conclusion about Answers Of Specific Heat. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Answers Of Specific Heat.

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Table of Contents Answers Of Specific Heat

1. Understanding the eBook Answers Of Specific Heat

- The Rise of Digital Reading Answers Of Specific Heat
- Advantages of eBooks Over Traditional Books
- 2. Identifying Answers Of Specific Heat
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Of Specific Heat
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answers Of Specific Heat
 - Personalized Recommendations
 - Answers Of Specific Heat User Reviews and Ratings
 - Answers Of Specific Heat and Bestseller Lists
- 5. Accessing Answers Of Specific Heat Free and Paid eBooks
 - Answers Of Specific Heat Public Domain eBooks
 - Answers Of Specific Heat eBook Subscription Services
 - Answers Of Specific Heat Budget-Friendly Options
- 6. Navigating Answers Of Specific Heat eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Of Specific Heat Compatibility with Devices
 - Answers Of Specific Heat Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Of Specific Heat
 - Highlighting and Note-Taking Answers Of Specific Heat
 - Interactive Elements Answers Of Specific Heat
- 8. Staying Engaged with Answers Of Specific Heat
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Of Specific Heat

9. Balancing eBooks and Physical Books Answers Of Specific Heat
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Of Specific Heat
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers Of Specific Heat
 - Setting Reading Goals Answers Of Specific Heat
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers Of Specific Heat
 - Fact-Checking eBook Content of Answers Of Specific Heat
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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