

Gravimetric Analysis of Chloride in Solution

Experiment 10
Analytical Chemistry
Fall 2023

Introduction

This experiment is designed to determine the concentration of chloride ions in a solution using gravimetric analysis. The method involves the precipitation of chloride ions as silver chloride (AgCl) using a known amount of silver ions. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is designed to illustrate the principles of gravimetric analysis and the importance of accurate weighing and drying of the precipitate. The experiment is designed to be performed in a laboratory setting and is suitable for students with a basic understanding of chemistry.

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Procedure

The procedure involves the precipitation of chloride ions as silver chloride (AgCl) using a known amount of silver ions. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution.

Data Analysis

Sample #	Sample #	Sample #
Mass of sample	0.250 g	0.250 g
Mass of precipitate	0.450 g	0.450 g
Mass of chloride ion	0.150 g	0.150 g

The mass of chloride ion in the sample is calculated as follows:

$$\text{Mass of chloride ion} = \text{Mass of precipitate} \times \frac{\text{Molar mass of Cl}^-}{\text{Molar mass of AgCl}}$$

The mass of chloride ion in the sample is calculated as follows:

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WJ Hussar



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