

Oxidation-Reduction Worksheet

For each reaction below, identify the atom oxidized, the atom reduced, the oxidizing agent, the reducing agent, the oxidation half reaction, the reduction half reaction, and then balance the equation by the method of oxidation-reduction showing all electrons transfers.

- $$\overset{0}{\text{Mg}} + 2\overset{+1}{\text{H}}\overset{-1}{\text{Cl}} \rightarrow \overset{+2}{\text{Mg}}\overset{-2}{\text{Cl}}_2 + \overset{0}{\text{H}}_2$$

$$\overset{0}{\text{Mg}} \xrightarrow{-2e^-} \overset{+2}{\text{Mg}} \quad 2\overset{+1}{\text{H}} \xrightarrow{+2e^-} \overset{0}{\text{H}}_2$$
- $$\overset{0}{2\text{Fe}} + 3\overset{+3}{\text{V}}_2\overset{-2}{\text{O}}_3 \rightarrow \overset{+3}{\text{Fe}}_2\overset{-2}{\text{O}}_3 + 6\overset{+2}{\text{V}}\overset{-2}{\text{O}}$$

$$2\overset{0}{\text{Fe}} \xrightarrow{-6e^-} \overset{+3}{\text{Fe}}_2 \quad 3\overset{+3}{\text{V}}_2 \xrightarrow{+6e^-} 6\overset{+2}{\text{V}}$$
- $$\overset{+7}{\text{KMnO}}_4 + \overset{+5}{\text{KNO}}_2 + 2\overset{+2}{\text{H}}_2\overset{-2}{\text{SO}}_4 \rightarrow \overset{+2}{\text{Mn}}\overset{-2}{\text{SO}}_4 + 2\overset{+2}{\text{H}}_2\text{O} + \overset{+5}{\text{KNO}}_3 + \overset{-2}{\text{K}_2\text{SO}}_4$$
- $$\text{K}_2\text{Cr}_2\text{O}_7 + \text{SnCl}_2 + \text{HCl} \rightarrow \text{CrCl}_3 + \text{SnCl}_4 + \text{KCl} + \text{H}_2\text{O}$$
- $$\text{KMnO}_4 + \text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{Cl}_2 + \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$$
- $$\text{K}_2\text{Cr}_2\text{O}_7 + \text{H}_2\text{O} + \text{S} \rightarrow \text{SO}_2 + \text{KOH} + \text{Cr}_2\text{O}_3$$
- $$\text{KClO}_3 + \text{C}_{12}\text{H}_{22}\text{O}_{11} \rightarrow \text{KCl} + \text{H}_2\text{O} + \text{CO}_2$$
- $$\text{H}_2\text{C}_2\text{O}_4 + \text{K}_2\text{MnO}_4 \rightarrow \text{CO}_2 + \text{K}_2\text{O} + \text{Mn}_2\text{O}_3 + \text{H}_2\text{O}$$
- $$\text{Mn}(\text{NO}_3)_2 + \text{NaBiO}_3 + \text{HNO}_3 \rightarrow \text{HMnO}_4 + \text{Bi}(\text{NO}_3)_3 + \text{NaNO}_3 + \text{H}_2\text{O}$$
- $$\text{H}_2\text{C}_2\text{O}_4 + \text{KMnO}_4 \rightarrow \text{CO}_2 + \text{K}_2\text{O} + \text{Mn}_2\text{O}_3 + \text{H}_2\text{O}$$

Analyzing Oxidation Reduction Reactions Answer Key

Dr. Alok kumar Dash

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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