

Pre-lab Questions for Experiment #5

Rates of Chemical Reactions I: A Clock Reaction

1) Complete and balance the following reactions:



2) The rate law for the reaction



is $\text{rate} = k[\text{H}_2][\text{NO}]^2$. Which of the following mechanisms can be ruled out on the basis of the observed rate expression?

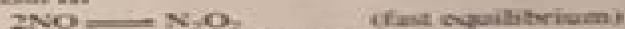
Mechanism I



Mechanism II



Mechanism III



3) Consider the reaction



From the following data obtained at a certain temperature, determine the order of the reaction and calculate the rate constant:

[A] (M)	[B] (M)	Rate (M/s)
1.50	1.50	3.20×10^{-1}
1.50	2.50	3.20×10^{-1}
3.00	1.50	6.40×10^{-1}

4) The following gas-phase reaction was studied at 290°C by observing the change in pressure as a function of time in a constant-volume vessel:



Determine the order of the reaction and the rate constant based on the following data:

Time (s)	Partial Pressure of $\text{ClCO}_2\text{CCl}_3$ (mmHg)
0	15.76
181	12.64
513	8.73
1164	4.44

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Hunt



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