

[illegible]

2.50 ml of  
1.0 M NaOH

**Abstract**

- ### ५. मरु-संस्कृत संज्ञा

**☞** কোলাস পানিসূত্র লবণের দ্রাব্যতা আপনাত্মা কৃষ্ণির সাথে সাথে এক পর্যায়ে কৃষ্ণি না পেয়ে হ্রাস পায়। এর কারণ, আপনাত্মা কৃষ্ণির সাথে সাথে দ্রাব্যতা কৃষ্ণি পেতে থাকলেও একটি নির্দিষ্ট আপনাত্মার কোলাস পানি দ্রুত লবণ নিষ্কৃতিত হয়। তখন দ্রাব্যতা হ্রাস পেতে থাকে। উদাহরণস্বরূপ  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর কথা বলার যায়।  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  এর ক্ষেত্রে পানিতে দ্রবীভূত হওয়ার সময় বিয়োজন প্রক্রিয়াটি তাপস্বরী বিক্রিয়া কিন্তু  $32.4^\circ\text{C}$  আপনাত্মার উপরে  $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}$  যখন  $\text{Na}_2\text{SO}_4$  এ পরিণত হয় তখন বিয়োজন প্রক্রিয়াটি তাপোৎপাদী হয়। এ কারণে  $32.4^\circ\text{C}$  আপনাত্মার উপরে আপনাত্মা কৃষ্ণিজনিত কারণে দ্রাব্যতা হ্রাস ঘটে।

w = SMV  
 =  $(0.5 \times 98 \times 0.25) \text{ g}$   
 = 12.25 g  
 অতএব, শুষ্কত H<sub>2</sub>SO<sub>4</sub> এর ভর  
 = 12.25 g

উদাহরণের A-পাত্রে অবস্থিত,  
 $\text{H}_2\text{SO}_4$  এর আণবিক ভর,  
 $M = 1 \times 2 + 32 + 16 \times 4$   
 $= 2 + 32 + 64$   
 $= 98 \text{ g/mol}$   
 দ্রবণের ঘনমাত্রা,  
 $S = 0.5 \text{ M}$   
 দ্রবণের আয়তন,  
 $V = 250 \text{ mL}$   
 $= \frac{250}{1000} = 0.25 \text{ L}$

$$\text{[3]} \quad \text{HNO}_3 \rightarrow \text{H}_2\text{SO}_4 \text{ की मात्रा ज्ञात करें} = \left( 0.5 \times \frac{250}{1000} \right) \text{ mol}$$

$$= 0.125 \text{ mol}$$

பொருள்-11:  $\text{Na}_2\text{CO}_3$  க்கு கிடைக்கும் அளவு =  $\left(0.11 \times \frac{200}{1000}\right) \text{ mol}$   
 $= 0.022 \text{ mol}$

अण्विहित समीकरण:  $\text{H}_2\text{SO}_4 + \text{Na}_2\text{CO}_3 \rightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$   
अधिकतम शुद्ध भार, 1 mol  $\text{H}_2\text{SO}_4 = 1 \text{ mol Na}_2\text{CO}_3$

1 mol  $\text{Na}_2\text{CO}_3$  বিক্রিয়া করে  $\text{H}_2\text{SO}_4$  এর 1 mol এর সাথে  
 0.077 mol " " " " (0.077) mol "

অনশিট  $\text{H}_2\text{SO}_4$  এর পরিমাণ =  $(0.125 - 0.077) \text{ mol}$   
 =  $0.048 \text{ mol}$

পাত্র-A এর দ্রবণের সাথে পাত্র-B এর দ্রবণের বিক্রিয়ার ফলে ০.৫ mol  $H_2SO_4$  অবশিষ্ট থাকায় মিশ্রিত দ্রবণের প্রকৃতি হবে অম্লিক।

**উদাহরণ ২** একজন ছাত্র ল্যাবরেটরিতে X-পদার্থিতে 10 mg এর ত্রুটি পরিমাণ এবং Y-পদার্থিতে 0.5 g পরিমাণ 2%  $\frac{W}{V}$  HCl এর দ্রবণ তৈরি করে।

**Abstract**

- নোড কাকে বলে?
- $Cl_2$  অণুজার— ব্যাখ্যা করো।
- উদ্ভীপকের এলিভটির মৌল্যত্ব ঘনমাত্রা নির্ণয় করো।
- উদ্ভীপকের কোন পদ্ধতি পরিবেশবান্ধব? বিশ্লেষণ করো।

२. नए व्यंजन का चिह्न

**উদাহরণ ১:** নিউক্লিয়াসের চারদিকে যে নির্দিষ্ট ত্রিমাত্রিক স্থানে কোন ইলেকট্রন পাওয়ার সম্ভাব্যতা শূন্য তাকে নোড বলে।

১৩। সমযোজী বন্ধনে আবদ্ধ উভয় পরমাণুর তড়িৎ ঋণাত্মকতা সম হলে অর্থাৎ তড়িৎ ঋণাত্মকতার পার্থক্য শূন্য হলে সমযোজী বন্ধন বিশুদ্ধ ও অপোলার হয়। আর তড়িৎ ঋণাত্মকতার পার্থক্য 0.5 এর কম হলে অপোলার হয়। যেহেতু  $Cl_1$  এর তড়িৎ ঋণাত্মকতা 3.0 ও  $C_2$  অণুতে তড়িৎ ঋণাত্মকতার পার্থক্য  $(3 - 3) = 0$ , তাই  $Cl_2$  বিশুদ্ধ সমযোজী ও অপোলার যৌগ। উল্লিখিত মৌলিক অণুর ক্ষেত্রে এছাড়াও থাকে।

গ) উদ্দীপকে বর্ণিত এসিডটি হলো  $2\% \left(\frac{W}{V}\right) HCl$

2%  $\left(\frac{W}{V}\right)$  HCl এর অর্থ হলো—

100 mL. મુશળ HCl એનું ઘન 2g

$$\therefore 1000 \text{ mL} = \left( \frac{2 \times 1000}{100} \right) \text{ g} = 20 \text{ g}$$

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**JA Banks**



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## **Table of Contents Answer Of Hsc Chemistry 1st Paper 15 Ctg Board**

1. Understanding the eBook Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
  - The Rise of Digital Reading Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
  - Advantages of eBooks Over Traditional Books
2. Identifying Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
  - Personalized Recommendations
  - Answer Of Hsc Chemistry 1st Paper 15 Ctg Board User Reviews and Ratings
  - Answer Of Hsc Chemistry 1st Paper 15 Ctg Board and Bestseller Lists
5. Accessing Answer Of Hsc Chemistry 1st Paper 15 Ctg Board Free and Paid eBooks
  - Answer Of Hsc Chemistry 1st Paper 15 Ctg Board Public Domain eBooks
  - Answer Of Hsc Chemistry 1st Paper 15 Ctg Board eBook Subscription Services
  - Answer Of Hsc Chemistry 1st Paper 15 Ctg Board Budget-Friendly Options
6. Navigating Answer Of Hsc Chemistry 1st Paper 15 Ctg Board eBook Formats

- ePub, PDF, MOBI, and More
  - Answer Of Hsc Chemistry 1st Paper 15 Ctg Board Compatibility with Devices
  - Answer Of Hsc Chemistry 1st Paper 15 Ctg Board Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
  - Highlighting and Note-Taking Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
  - Interactive Elements Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
8. Staying Engaged with Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
9. Balancing eBooks and Physical Books Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
- Setting Reading Goals Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
- Fact-Checking eBook Content of Answer Of Hsc Chemistry 1st Paper 15 Ctg Board
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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