#### Post-Lab Questions

- 1. From the experiment, which reactant is the limiting reactant? Which one is the reactant in excess? How do you determine which one is the limiting reactant?
- 2. If you weigh the watch glass then write your name on it with the wax pencil, how will that affect your percent yield?
- 3. What factors could cause the percent yield of a product to be less than 100%?
- 4. What factors could cause the percent yield of a product to be more than 100%?
- Balance the following reaction: \_\_\_C<sub>3</sub>H<sub>8</sub> + \_\_\_O<sub>2</sub> → \_\_\_CO<sub>2</sub> + \_\_\_H<sub>2</sub>O
  - a. 4.30 g of C<sub>3</sub>H<sub>8</sub> reacts with 4.30 g of O<sub>2</sub>. Which one is the limiting reagent, and which one is in excess? Explain your reasoning.
  - b. Based on the limiting reagent, what are the theoretical yields of CO<sub>2</sub> and H<sub>2</sub>O?
  - c. You perform the experiment and produce 2.67 g of CO<sub>2</sub> and 0.79 g of H<sub>2</sub>O. What are your percent yields for each product?

# **Answers Key To Chemistry 1211 Lab Manual**

J Ma

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