



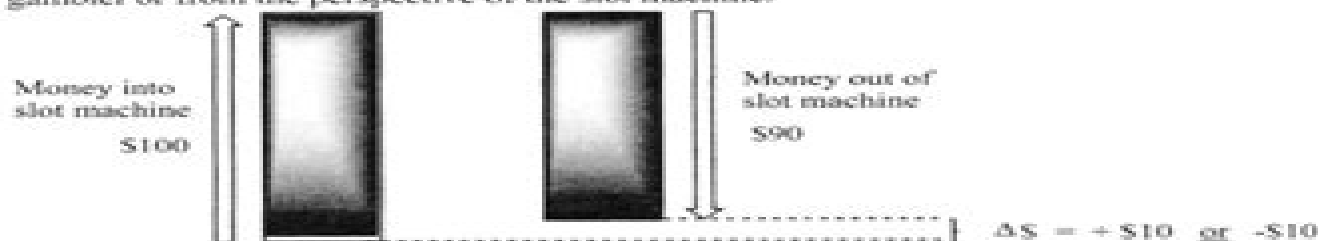
ChemActivity 6

Part A: Energy Diagrams

(What is the sign of $\Delta H_{\text{reaction}}$ for an exothermic reaction?)

Model 1A: The Slot Machine

Imagine a gambler sitting at a slot machine. Over the course of an hour the gambler puts \$100 into the machine and the machine pays out \$90. We know a net change (Δ) of ten dollars occurred, but what is the sign of this change in money (ΔS). To determine if $\Delta S = +10$ or -10 we need to decide if we calculate ΔS from the perspective of the gambler or from the perspective of the slot machine.



To translate this analogy into chemistry remember that:

- **Money = energy.**
- **The slot machine = the reaction.**
- **The gambler = the experimenter (you).**
- **All changes (Δ) will be calculated from the perspective of the reaction (a.k.a. the slot machine).**

Critical Thinking Questions

1. Using the convention listed above for calculating changes (Δ), what is the sign of ΔS in the example in Model 1A?
2. Based on this same convention, what sign must be attached to the \$100 that the gambler put into the slot machine in Model 1A? The sign of the \$90 that comes out is the opposite. **Add the appropriate sign to each value in Model 1A.**
3. On a different day the gambler puts in a total of \$20 and the slot machine pays out a total of \$100. What is the sign and value of ΔS for this day?



Answers Chemactivity 6

Michael P. Garoutte

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1. Understanding the eBook Answers Chemactivity 6

- The Rise of Digital Reading Answers Chemactivity 6
- Advantages of eBooks Over Traditional Books
- 2. Identifying Answers Chemactivity 6
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Chemactivity 6
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answers Chemactivity 6
 - Personalized Recommendations
 - Answers Chemactivity 6 User Reviews and Ratings
 - Answers Chemactivity 6 and Bestseller Lists
- 5. Accessing Answers Chemactivity 6 Free and Paid eBooks
 - Answers Chemactivity 6 Public Domain eBooks
 - Answers Chemactivity 6 eBook Subscription Services
 - Answers Chemactivity 6 Budget-Friendly Options
- 6. Navigating Answers Chemactivity 6 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Chemactivity 6 Compatibility with Devices
 - Answers Chemactivity 6 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Chemactivity 6
 - Highlighting and Note-Taking Answers Chemactivity 6
 - Interactive Elements Answers Chemactivity 6
- 8. Staying Engaged with Answers Chemactivity 6
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Chemactivity 6

9. Balancing eBooks and Physical Books Answers Chemactivity 6
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Chemactivity 6
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers Chemactivity 6
 - Setting Reading Goals Answers Chemactivity 6
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers Chemactivity 6
 - Fact-Checking eBook Content of Answers Chemactivity 6
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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