

# 2.5

## ELEMENTS AND COMPOUNDS

### Section Review

#### Objectives

1. Explain the difference between elements and a compound
2. Distinguish between a substance and a mixture
3. Identify the chemical symbols for elements and write a chemical formula for a compound

#### Vocabulary

- |              |                |
|--------------|----------------|
| 1. element   | 2. compound    |
| 3. substance | 4. homogeneous |

#### Part A Completion

Fill in the blank with the word or phrase that best completes or describes each statement. Write your answer in the space provided.

- |                                                                                                                                                                    |                                                                                                                                                                    |
|--------------------------------------------------------------------------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------------------------------------------------------------------------|
| 1. An element is a substance that cannot be broken down into simpler substances by chemical means.                                                                 | 2. A compound is a substance that is made up of two or more different elements.                                                                                    |
| 3. A mixture is a combination of two or more substances that are not chemically combined.                                                                          | 4. A homogeneous mixture is a mixture in which the components are uniformly distributed.                                                                           |
| 5. A heterogeneous mixture is a mixture in which the components are not uniformly distributed.                                                                     | 6. A pure substance is a substance that has a fixed composition and distinct chemical and physical properties.                                                     |
| 7. A chemical formula is a symbolic representation of a chemical compound.                                                                                         | 8. A chemical reaction is a process in which one or more substances are converted into one or more different substances.                                           |
| 9. A chemical equation is a symbolic representation of a chemical reaction.                                                                                        | 10. A balanced chemical equation is a chemical equation in which the number of atoms of each element is the same on both sides.                                    |
| 11. A chemical reaction can be identified by the formation of a precipitate, the evolution of a gas, the formation of a color change, or the formation of a solid. | 12. A chemical reaction can be identified by the formation of a precipitate, the evolution of a gas, the formation of a color change, or the formation of a solid. |
| 13. A chemical reaction can be identified by the formation of a precipitate, the evolution of a gas, the formation of a color change, or the formation of a solid. | 14. A chemical reaction can be identified by the formation of a precipitate, the evolution of a gas, the formation of a color change, or the formation of a solid. |
| 15. A chemical reaction can be identified by the formation of a precipitate, the evolution of a gas, the formation of a color change, or the formation of a solid. | 16. A chemical reaction can be identified by the formation of a precipitate, the evolution of a gas, the formation of a color change, or the formation of a solid. |

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**Rex M. Heyworth**



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