

1999 Advanced Placement Chemistry Exam

Section I: Multiple Choice

Time — 1 hour and 30 minutes

NO CALCULATORS MAY BE USED WITH SECTION I.

Note: For all questions, assume that the temperature is 298 K, the pressure is 1.00 atmosphere, and solutions are aqueous unless otherwise specified.

Throughout the test the following symbols have the definitions specified unless otherwise noted.

T = temperature
 P = pressure
 V = volume
 S = entropy
 H = enthalpy
 G = free energy
 R = molar gas constant
 n = number of moles
 M = molar

m = molal
 L, mL = liter(s), milliliter(s)
 g = gram(s)
 nm = nanometer(s)
 atm = atmosphere(s)
 J, kJ = joule(s), kilojoule(s)
 V = volt(s)
 mol = mole(s)

Part A

Directions: Each set of lettered choices below refers to the numbered statements immediately following it. Select the one lettered choice that best fits each statement and then fill in the corresponding oval on the answer sheet. A choice may be used once, more than once, or not at all in each set.

Questions 1–4 refer to the following types of energy.

- (A) Activation energy
- (B) Free energy
- (C) Ionization energy
- (D) Kinetic energy
- (E) Lattice energy

1. The energy required to convert a ground-state atom in the gas phase to a gaseous positive ion
2. The energy change that occurs in the conversion of an ionic solid to widely separated gaseous ions
3. The energy in a chemical or physical change that is available to do useful work
4. The energy required to form the transition state in a chemical reaction

Questions 5–8 refer to atoms for which the occupied atomic orbitals are shown below.

- (A) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (B) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}}$
- (C) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (D) $1s \underline{\hspace{1cm}} 2s \underline{\hspace{1cm}} 2p \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$
- (E) $[Ar] 4s \underline{\hspace{1cm}} 3d \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}} \underline{\hspace{1cm}}$

5. Represents an atom that is chemically unreactive
6. Represents an atom in an excited state
7. Represents an atom that has four valence electrons
8. Represents an atom of a transition metal

Questions 9–12 refer to aqueous solutions containing 1:1 mole ratios of the following pairs of substances. Assume all concentrations are 1 M .

2008 Ap Chemistry Exam Multiple Choice Answers

Kathleen Armour



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Executive Thinking Kevin Baize OD,Morris Graham PhD,2011-02-18 Life is about how much we think Thinking is about how much mental capacity we possess Capacity in addition to our abilities and conscientiousness is about how much we can process combinations of verbal height quantitative width and spatial depth with decisiveness direction and speed No matter where we go or what we do as executives we take our thinking with us That may spoil everything because to a great extent we do and accomplish what we think about Our thoughts mold our aspirations attitudes and accomplishments during our life In other words our careers and lives are influenced more by the power of our thoughts than anything else The bad news is that most of us never fully use our mental capacities and never achieve our potential The good news is that neural technologies are now available to transform our thinking into the higher realms of brilliance Developing the spatial capacity to think higher wider and deeper means breaking away from the effects of years of flat thinking or educational backgrounds that stifles creative innovative potential Expand your mental agility through a development of higher order processes and discover a whole new world mentally in Executive Thinking **Guide to College Visits** Princeton Review,2007-04-10 A practical handbook covers everything needed to plan a successful visit to campuses across the country with information on nearby accommodations for every budget the best and worst times to visit itineraries and campus tours **Excel With Objective Questions In Chemistry** Prof. S. K. Khanna,Dr. N. K. Verma,Dr. B. Kapila,2006

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