

5-2 Review and Reinforcement Part II

Quantum Theory and the Atom

Complete each statement, using the vocabulary and the characteristics and history of quantum theory.

characteristic

Pauli exclusion principle

quantum-mechanical model

quantum-mechanical

electron density

orbital

principal energy levels

_____ orbital _____ 1-4, proposed to explain around the nucleus where electrons in atoms are farthest from.

_____ electron density _____ 2, the density of an electron cloud

_____ quantum-mechanical _____ 3, involves developing a quantum-mechanical model for atoms

Pauli exclusion principle _____ 4, states that any two electrons cannot occupy the same orbital, and that those

_____ quantum-mechanical model approach spins

_____ quantum-mechanical model _____ 5, explains the properties of atoms by treating the electrons as waves and

_____ quantum-mechanical model

_____ quantum-mechanical model _____ 6, the model of atoms that involves

_____ electron spin _____ 7, the introduction of quantum-mechanical model of atoms

Complete each statement, using the history of the quantum-mechanical model of atoms.

_____ 1. _____ 1. The electron cloud is _____ when the probability of finding an electron is

- highest
- lowest
- highest density
- lowest density

_____ 2. _____ 2. The low principal energy level of the hydrogen atom contains only one

- orbital
- or orbitals
- or orbitals
- orbitals

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Gerard Kraus



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