

Ch. 9 Worksheet #1: Molar Ratio Practice Problems

Following each equation are two requests for molar ratios from the equation.



N_2 to H_2 : 1:3 NH_3 to H_2 : 2:3



O_2 to SO_3 : 1:2 O_2 to SO_2 : 1:2



PCl_3 to Cl_2 : 1:1 PCl_3 to PCl_5 : 1:1



NH_3 to N_2 : 4:2 $\rightarrow$ 2:1 H_2O to O_2 : 6:3 $\rightarrow$ 2:1



CO to CO_2 : 3:3 $\rightarrow$ 1:1 Fe to CO : 2:3

Mole to Mole Practice Problems

Here's the equation to use for all three problems:



1) How many moles of H_2O are produced when 5.00 moles of oxygen are used?

Remember: *Starting with five moles of oxygen and based on the balanced equation, for every 1 mole of oxygen used, two moles of water are produced.*

$$\underline{5.00} \text{ moles O}_2 \times \frac{2 \text{ mole H}_2\text{O}}{1 \text{ mole O}_2} = \underline{10} \text{ moles H}_2\text{O}$$

2) If 3.00 moles of H_2O are produced, how many moles of oxygen must be consumed?

$$3.00 \text{ moles of H}_2\text{O} \times 1 \text{ mole O}_2 / 2 \text{ mole H}_2\text{O} = 1.5 \text{ moles O}_2$$

3) Using the data from problem two, how many moles of hydrogen gas must be used,?

$$3.00 \text{ moles of H}_2\text{O} \times 2 \text{ mole H}_2 / 2 \text{ mole H}_2\text{O} = 3.00 \text{ mole of H}_2$$

4) Suppose 4.00 grams of H_2 were used? How many grams of water would be produced?

$$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O} \quad 2(1.008) = 2.016$$

$$2(1.008) + 16 = 18 \text{ g/mol}$$

$$2(18) = 36 \text{ g of H}_2\text{O}$$

$$4\text{g} / 2.016 \text{ g/mol} = 1.98 \text{ mol} \rightarrow 2 \text{ mol}$$

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