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assuming CeCl_2 CuCl_2 CuCl_2 is available in excess

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 $\text{Br}_2 + \text{C}_6\text{H}_{10} \rightarrow \text{C}_6\text{H}_8 + \text{H}_2\text{O}$
 $\text{H}_2\text{O} + \text{KBr} + \text{CO}_2 \rightarrow \text{K}_2\text{SO}_3 + \text{Ga}_2\text{SO}_3 + \text{NaBr}$
 $\text{NaOH} + \text{H}_2\text{SO}_4 \rightarrow \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$
 br using the following equation
 $\text{Br}_2 + \text{C}_6\text{H}_{10} \rightarrow \text{C}_6\text{H}_8 + \text{H}_2\text{O}$

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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
 $\text{CaH}_2 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
 $\text{HBr} + \text{KHCO}_3 \rightarrow \text{H}_2\text{O} + \text{KBr} + \text{CO}_2$
 $\text{GaBr}_3 + \text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + \text{NaBr}$
 SnF_2

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 $\text{SiO}_2 + 3\text{C} \rightarrow \text{SiC} + 2\text{CO}$

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 $\text{N}_2 + \text{F}_2 \rightarrow \text{NF}_3$
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 $\text{KBr} + \text{GaBr}_3 + \text{Na}_2\text{SO}_3 \rightarrow \text{Ga}_2\text{SO}_3 + \text{NaBr} + \text{CO}_2$
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same thing goes with the o3 i am so lost

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web a common type of stoichiometric relationship is the mole ratio which relates the amounts in moles of any two substances in a chemical reaction we can write a mole ratio for a pair of substances by looking at the coefficients in front of each species in

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