

Gravimetric Analysis of Chloride in Solution

Experiment 10
Lab 10
Date: _____

Introduction

This experiment involves the gravimetric analysis of a sample solution containing chloride ions. The sample solution is a mixture of sodium chloride and sodium nitrate. The sample is weighed and then dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride in the original sample. The experiment is performed in two parts. In the first part, the sample is weighed and dissolved. In the second part, the sample is treated with silver nitrate and the precipitate is filtered, dried, and weighed.

The sample is weighed and dissolved in water. The solution is then treated with silver nitrate, which reacts with the chloride ions to form a precipitate of silver chloride. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride in the original sample.

Procedure

The procedure for this experiment is as follows: 1. Weigh a sample of the solution. 2. Dissolve the sample in water. 3. Add silver nitrate to the solution. 4. Filter the precipitate. 5. Dry the precipitate. 6. Weigh the precipitate.

Data Analysis

	Sample 1	Sample 2
Mass of sample	0.500 g	0.500 g
Mass of precipitate	0.850 g	0.850 g
Mass of chloride in sample	0.350 g	0.350 g

The mass of chloride in the sample is calculated as follows: $m_{\text{Cl}^-} = m_{\text{precipitate}} \times \frac{m_{\text{Cl}^-}}{m_{\text{AgCl}}}$

For Sample 1: $m_{\text{Cl}^-} = 0.850 \text{ g} \times \frac{35.45 \text{ g}}{143.32 \text{ g}} = 0.350 \text{ g}$

For Sample 2: $m_{\text{Cl}^-} = 0.850 \text{ g} \times \frac{35.45 \text{ g}}{143.32 \text{ g}} = 0.350 \text{ g}$

Analysis Of Chloride Lab Report

SJ Ball



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Table of Contents Analysis Of Chloride Lab Report

1. Understanding the eBook Analysis Of Chloride Lab Report
 - The Rise of Digital Reading Analysis Of Chloride Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Analysis Of Chloride Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analysis Of Chloride Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Analysis Of Chloride Lab Report
 - Personalized Recommendations
 - Analysis Of Chloride Lab Report User Reviews and Ratings
 - Analysis Of Chloride Lab Report and Bestseller Lists

5. Accessing Analysis Of Chloride Lab Report Free and Paid eBooks
 - Analysis Of Chloride Lab Report Public Domain eBooks
 - Analysis Of Chloride Lab Report eBook Subscription Services
 - Analysis Of Chloride Lab Report Budget-Friendly Options
6. Navigating Analysis Of Chloride Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Analysis Of Chloride Lab Report Compatibility with Devices
 - Analysis Of Chloride Lab Report Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analysis Of Chloride Lab Report
 - Highlighting and Note-Taking Analysis Of Chloride Lab Report
 - Interactive Elements Analysis Of Chloride Lab Report
8. Staying Engaged with Analysis Of Chloride Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analysis Of Chloride Lab Report
9. Balancing eBooks and Physical Books Analysis Of Chloride Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analysis Of Chloride Lab Report
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Analysis Of Chloride Lab Report
 - Setting Reading Goals Analysis Of Chloride Lab Report
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Analysis Of Chloride Lab Report
 - Fact-Checking eBook Content of Analysis Of Chloride Lab Report
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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