Name: Key

For each reaction:

- * Assign Oxidation Numbers
- * Write an Oxidation and Reduction Half Reaction
- *Circle the Oxidizing Agent and Underline the Reducing Agent.

1) Fe⁻³ + 5n → Fe⁻² + 5n⁻²

Oxidation: Snº -> Snº + 2e

Reduction: (Fe') . Is' -> Fe'

2) AlF, * Bo -> Al * BoF.

Oxidation: Bot + Bort + 2e-

Reduction: (Al-3) + 3e - Al-

3) H₂O₂ + H₂SO₂ → H₂O + H₂SO₂

Oxidation: 5-1 + 5-1 + 2e

Reduction: (O1) + Ie: + O-

4) $H_1SO_1 + KMnO_4 \rightarrow MnSO_4 + K_1SO_4 + H_2O_5$

Oxidation: 5" + 5" + 2e

Reduction: (Mn-) + 5e - Mn-2

5) K2Cr2O7 + H2C2O4 + H2SO4 + K2SO4 + Cr2(SO4)3 + CO2 + H2O

Oxidation: C1 -> C-1 + 1e

Reduction: (Cris) . 3c -> Cris

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Paul Arthur, Otto Mitchell Smith

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Chemical Engineering Progress ,1967

Stanitski, 1986

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