

▼ Part A

If a reaction mixture initially contains a CO concentration of 0.1510 M and a Cl_2 concentration of 0.175 M at 1000 K . What is the equilibrium concentration of CO at 1000 K ?

Express your answer in molarity to three significant figures.

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$[\text{CO}] =$ M

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✖ Incorrect; Try Again; 2 attempts remaining

It appears that you have not included the equilibrium constant in your calculation. Recall that the equilibrium concentration of CO is related to the equilibrium constant and the change in concentrations of the product and reactants:

$$K_c = \frac{[\text{COCl}_2]}{[\text{CO}][\text{Cl}_2]}$$

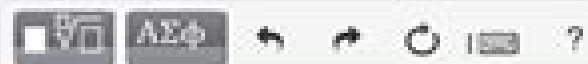
You may want to review [Hint 1. Prepare an ICE table.](#)

▼ Part B

What is the equilibrium concentration of Cl_2 at 1000 K ?

Express your answer in molarity to three significant figures.

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$[\text{Cl}_2] =$ M

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Table of Contents 2neutralization Reaction Section Review

1. Understanding the eBook 2neutralization Reaction Section Review
 - The Rise of Digital Reading 2neutralization Reaction Section Review
 - Advantages of eBooks Over Traditional Books
2. Identifying 2neutralization Reaction Section Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2neutralization Reaction Section Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2neutralization Reaction Section Review
 - Personalized Recommendations
 - 2neutralization Reaction Section Review User Reviews and Ratings
 - 2neutralization Reaction Section Review and Bestseller Lists
5. Accessing 2neutralization Reaction Section Review Free and Paid eBooks
 - 2neutralization Reaction Section Review Public Domain eBooks
 - 2neutralization Reaction Section Review eBook Subscription Services
 - 2neutralization Reaction Section Review Budget-Friendly Options

6. Navigating 2neutralization Reaction Section Review eBook Formats
 - ePub, PDF, MOBI, and More
 - 2neutralization Reaction Section Review Compatibility with Devices
 - 2neutralization Reaction Section Review Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 2neutralization Reaction Section Review
 - Highlighting and Note-Taking 2neutralization Reaction Section Review
 - Interactive Elements 2neutralization Reaction Section Review
8. Staying Engaged with 2neutralization Reaction Section Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 2neutralization Reaction Section Review
9. Balancing eBooks and Physical Books 2neutralization Reaction Section Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 2neutralization Reaction Section Review
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 2neutralization Reaction Section Review
 - Setting Reading Goals 2neutralization Reaction Section Review
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 2neutralization Reaction Section Review
 - Fact-Checking eBook Content of 2neutralization Reaction Section Review
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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