

CHAPTER 18 REVIEW

Ions in Aqueous Solutions and Colligative Properties

SECTION 1

SHORT ANSWER Answer the following questions in the space provided.

1. Give the formulas of *buttle 1* on page 677 of the answer packet for each of the following compounds in water:

- _____ a. magnesium nitrate
 _____ b. sodium sulfate
 _____ c. calcium sulfide
 _____ d. potassium phosphate

2. List each of magnesium nitrate in *buttle 1* in water:

- _____ a. Write the formula for magnesium nitrate.
 _____ b. How many moles of ions per mole of ions is there?
 _____ c. How many moles of ions are released into a solution made from 0.05 mol magnesium nitrate dissolved in water?

3. Write the formula for the precipitate formed:

- _____ a. when solutions of magnesium chloride and potassium phosphate are combined
 _____ b. when solutions of sodium sulfate and silver nitrate are combined

4. Write ionic equations for the dissolution of the following compounds:

- a. PbSO_4

 b. CaSO_4

5. a. Write the net ionic equation for the reaction that occurs when solutions of lead(II) nitrate and potassium sulfate are combined.

- b. What are the spectator ions in this reaction?

Aqueous Solution Section Review Answers

Rex M. Heyworth



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