

CHAPTER 10: Reactions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer on the line to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ for a reaction is 2076 kJ whereas a 4 mol/L of energy is released from the decomposition. Is it? _____
- _____ A compound has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If a 10 g sample undergoes a 1°C increase in temperature, how much energy is released? _____
- The product is a molecular form of oxygen, and the reactant has an enthalpy of 100 kJ/mol . _____
- What is the value of ΔH for this reaction? _____

Answers For Reaction Energy Section 1

**United States. Energy Research and
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