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AP CHEMISTRY MULTIPLE CHOICE QUESTIONS Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed
+ Doubled

(A) 95²

(B) 95

(C) (95)^{1/2}

(D) $\frac{1}{95}$

(E) $\frac{1}{(95)^2}$



$C_2H_2(g)$ is reduced by $H_2(g)$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $H_2(g)$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

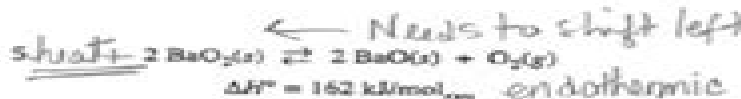
(A) III only

(B) I and II only

(C) I and III only

(D) II and III only

(E) I, II, and III



A sealed rigid vessel contains $BaO_2(s)$ in equilibrium with $BaO(s)$ and $O_2(g)$ as represented by the equation above. Which of the following changes will increase the amount of $BaO_2(s)$ in the vessel?

(A) Removing a small amount of $O_2(g)$ $R \times$

(B) Removing a small amount of $BaO(s)$ X

(C) Adding He gas to the vessel X

(D) Lowering the temperature L

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?

increase P°



shift towards fewer
+ make gas



When 4.00 mol each of $X(g)$ and $Y(g)$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $Z(g)$ is produced. What is the value of the equilibrium constant, K_c ?

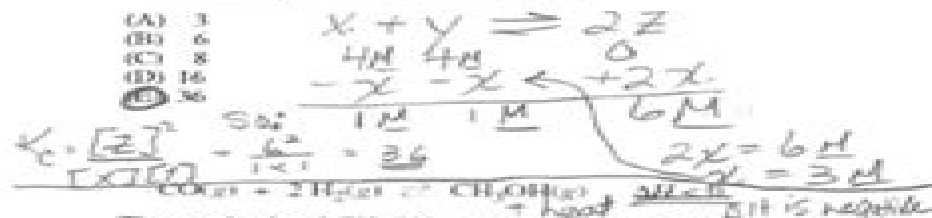
(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6. The synthesis of $CH_3OH(l)$ from $CO_2(g)$ and $H_2(g)$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease K_c

Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.

(B) K_c will increase because there are more reactant molecules than product molecules.

(C) K_c will decrease because the reaction is exothermic.

(D) K_c is constant and will not change.

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Gary S Thorpe



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Table of Contents 8 Ap Chemistry Multiple Choice Answers

1. Understanding the eBook 8 Ap Chemistry Multiple Choice Answers
 - The Rise of Digital Reading 8 Ap Chemistry Multiple Choice Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying 8 Ap Chemistry Multiple Choice Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 8 Ap Chemistry Multiple Choice Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from 8 Ap Chemistry Multiple Choice Answers
 - Personalized Recommendations
 - 8 Ap Chemistry Multiple Choice Answers User Reviews and Ratings
 - 8 Ap Chemistry Multiple Choice Answers and Bestseller Lists
5. Accessing 8 Ap Chemistry Multiple Choice Answers Free and Paid eBooks
 - 8 Ap Chemistry Multiple Choice Answers Public Domain eBooks

- 8 Ap Chemistry Multiple Choice Answers eBook Subscription Services
- 8 Ap Chemistry Multiple Choice Answers Budget-Friendly Options
- 6. Navigating 8 Ap Chemistry Multiple Choice Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - 8 Ap Chemistry Multiple Choice Answers Compatibility with Devices
 - 8 Ap Chemistry Multiple Choice Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 8 Ap Chemistry Multiple Choice Answers
 - Highlighting and Note-Taking 8 Ap Chemistry Multiple Choice Answers
 - Interactive Elements 8 Ap Chemistry Multiple Choice Answers
- 8. Staying Engaged with 8 Ap Chemistry Multiple Choice Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 8 Ap Chemistry Multiple Choice Answers
- 9. Balancing eBooks and Physical Books 8 Ap Chemistry Multiple Choice Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 8 Ap Chemistry Multiple Choice Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 8 Ap Chemistry Multiple Choice Answers
 - Setting Reading Goals 8 Ap Chemistry Multiple Choice Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 8 Ap Chemistry Multiple Choice Answers
 - Fact-Checking eBook Content of 8 Ap Chemistry Multiple Choice Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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