AP Scoring Guide Chemistry Unit 8 Progress Check MCQ Questions And Answer Keys

Buffer solutions containing Na₂CO₃ and NaHCO₃ range in pH from 10.0 to 11.0. The chemical equation below represents the equilibrium between CO₃²⁻ and H₂O, and the table lists the composition of four different buffer solutions at 25°C.

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W. C. SHOW			

Buffer	[NaHCO ₃]	[Na ₂ CO ₃]	PH
1	0.150	0,100	7
2	0.200	0.200	10.32
3	0.100	0.100	10.32
-4	0.100	0.200	190-

- Which of the following chemical equilibrium equations best shows what happens in the buffer solutions to minimize the change in pH when a small amount of a strong base is added?
- (A) $H_3O^+(aq) + OH^-(aq) \rightleftharpoons 2 H_3O(l)$
- B $HCO_3^-(aq) + OH^-(aq) \rightleftharpoons CO_3^{2-}(aq) + H_2O(l)$
- (c) $CO_3^{2-}(aq) + H_3O^+(aq) \rightleftharpoons HCO_3^-(aq) + H_2O(l)$
- $O_3^{2-}(aq) + H_2O(l) \rightleftharpoons HCO_3^{-}(aq) + OH^{-}(aq)$
- 2. Which of the following mathematical expressions can be used to determine the approximate pH of buffer 1

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Mary Millhollon, Richard H. Langley

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