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AP CHEMISTRY MULTIPLE CHOICE QUESTIONS

Unit 8 – Introduction to Equilibrium



The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



Reversed
+ Doubled

(A) 95²

(B) 95

(C) (95)^{1/2}(D) $\frac{1}{95}$ (E) $\frac{1}{(95)^2}$ 

$\text{C}_2\text{H}_4(\text{g})$ is reduced by $\text{H}_2(\text{g})$ in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- I. Changes in the partial pressure of $\text{H}_2(\text{g})$
- II. Changes in the particle size of the platinum catalyst
- III. Changes in the temperature of the reaction system

(A) III only

(B) I and II only

(C) I and III only

(D) II and III only

(E) I, II, and III



A sealed rigid vessel contains $\text{BaO}_2(\text{s})$ in equilibrium with $\text{BaO}(\text{s})$ and $\text{O}_2(\text{g})$ as represented by the equation above. Which of the following changes will increase the amount of $\text{BaO}_2(\text{s})$ in the vessel?

(A) Removing a small amount of $\text{O}_2(\text{g})$ $R \times$ (B) Removing a small amount of $\text{BaO}(\text{s})$ X (C) Adding He gas to the vessel X (D) Lowering the temperature L

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products?

increase P°



shift towards fewer
+ make gas



When 4.00 mol each of $\text{X}(\text{g})$ and $\text{Y}(\text{g})$ are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of $\text{Z}(\text{g})$ is produced. What is the value of the equilibrium constant, K_c ?

(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6. The synthesis of $\text{CH}_3\text{OH}(\text{g})$ from $\text{CO}(\text{g})$ and $\text{H}_2(\text{g})$ is represented by the equation above. The value of K_c for the reaction at 483 K is 14.5.

increase heat shifts left, decrease K_c

Which of the following explains the effect on the equilibrium constant, K_c , when the temperature of the reaction system is increased to 650 K?

(A) K_c will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.(B) K_c will increase because there are more reactant molecules than product molecules.(C) K_c will decrease because the reaction is exothermic.(D) K_c is constant and will not change.

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Ian Pickup



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