

Calculations in Analytical Chemistry

The Mole

The mole is the SI unit for amount of substance. It corresponds to the amount of a chemical substance that contains the same number of particles present in 12 grams of C-12 which is equivalent to Avogadro's number, that is, $N_A = 6.022 \times 10^{23}$.

The molar mass is given as the mass of a compound in grams per mole of that substance. It is calculated by adding the molar masses of all the component elements in the compound. Take glucose for example:

Element	Molar mass per element	Number of atoms	Mass
C	12 g/mol	6	72 g/mol
H	1 g/mol	12	12 g/mol
O	16 g/mol	6	96 g/mol
Molar Mass			180 g/mol

Mole and Millimole Calculations

The millimole is simply 1/1000 of a mole. It is sometimes used in calculations because of its convenience.

Example (Skoog, et al., 2014):

1. Find the number of moles and millimoles of benzoic acid (122.1 g/mol) that are contained in 2.00 g of pure acid.

If we use HBz to represent benzoic acid:

$$HBz_{mol} = 2.00 \text{ g}_{HBz} \times \frac{1 \text{ mol}_{HBz}}{122.1 \text{ g}_{HBz}} = 0.0164 \text{ mol}_{HBz}$$

To obtain the number of millimoles, the millimolar mass of benzoic acid will be used:

$$HBz_{mol} = 2.00 \text{ g}_{HBz} \times \frac{1 \text{ mmol}_{HBz}}{0.1221 \text{ g}_{HBz}} = 16.4 \text{ mmol}_{HBz}$$

2. What is the mass in grams of Na^+ (22.99 g/mol) in 25.0 g of Na_2SO_4 (142.0 g/mol)?

$$\begin{aligned} \text{mass}_{\text{Na}^+} &= 25.0 \text{ g}_{\text{Na}_2\text{SO}_4} \times \frac{1 \text{ mol}_{\text{Na}_2\text{SO}_4}}{142.0 \text{ g}_{\text{Na}_2\text{SO}_4}} \times \frac{2 \text{ mol}_{\text{Na}^+}}{1 \text{ mol}_{\text{Na}_2\text{SO}_4}} \times \frac{22.99 \text{ g}_{\text{Na}^+}}{1 \text{ mol}_{\text{Na}^+}} \\ \text{mass}_{\text{Na}^+} &= 8.10 \text{ g}_{\text{Na}^+} \end{aligned}$$

Analytical Chemistry 3

Michael Seilmaier

A red circular graphic with a gradient, appearing as a semi-circle or a stylized arrow pointing to the right, located on the right side of the slide.

Analytical Chemistry 3:

The book delves into Analytical Chemistry 3. Analytical Chemistry 3 is an essential topic that needs to be grasped by everyone, from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Analytical Chemistry 3, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to Analytical Chemistry 3
- Chapter 2: Essential Elements of Analytical Chemistry 3
- Chapter 3: Analytical Chemistry 3 in Everyday Life
- Chapter 4: Analytical Chemistry 3 in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of Analytical Chemistry 3. The first chapter will explore what Analytical Chemistry 3 is, why Analytical Chemistry 3 is vital, and how to effectively learn about Analytical Chemistry 3.
3. In chapter 2, the author will delve into the foundational concepts of Analytical Chemistry 3. The second chapter will elucidate the essential principles that need to be understood to grasp Analytical Chemistry 3 in its entirety.
4. In chapter 3, this book will examine the practical applications of Analytical Chemistry 3 in daily life. The third chapter will showcase real-world examples of how Analytical Chemistry 3 can be effectively utilized in everyday scenarios.
5. In chapter 4, this book will scrutinize the relevance of Analytical Chemistry 3 in specific contexts. The fourth chapter will explore how Analytical Chemistry 3 is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Analytical Chemistry 3. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Analytical Chemistry 3.

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Table of Contents Analytical Chemistry 3

1. Understanding the eBook Analytical Chemistry 3

- The Rise of Digital Reading Analytical Chemistry 3
- Advantages of eBooks Over Traditional Books
- 2. Identifying Analytical Chemistry 3
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Analytical Chemistry 3
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Analytical Chemistry 3
 - Personalized Recommendations
 - Analytical Chemistry 3 User Reviews and Ratings
 - Analytical Chemistry 3 and Bestseller Lists
- 5. Accessing Analytical Chemistry 3 Free and Paid eBooks
 - Analytical Chemistry 3 Public Domain eBooks
 - Analytical Chemistry 3 eBook Subscription Services
 - Analytical Chemistry 3 Budget-Friendly Options
- 6. Navigating Analytical Chemistry 3 eBook Formats
 - ePub, PDF, MOBI, and More
 - Analytical Chemistry 3 Compatibility with Devices
 - Analytical Chemistry 3 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Analytical Chemistry 3
 - Highlighting and Note-Taking Analytical Chemistry 3
 - Interactive Elements Analytical Chemistry 3
- 8. Staying Engaged with Analytical Chemistry 3
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Analytical Chemistry 3

9. Balancing eBooks and Physical Books Analytical Chemistry 3
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Analytical Chemistry 3
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Analytical Chemistry 3
 - Setting Reading Goals Analytical Chemistry 3
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Analytical Chemistry 3
 - Fact-Checking eBook Content of Analytical Chemistry 3
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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