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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

bond	kJ mol ⁻¹	kJ mol ⁻¹
bond energy in kJ/mol	440	435

Use the bond energies in Table 3.1 and ΔH to calculate the bond energy of one N-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$\begin{aligned}
 & (500) + 3(440) \\
 & 440 + 3(440) \quad \quad \quad 1660 \text{ kJ}
 \end{aligned}$$
- Calculate the energy released when bonds form in the products.

$$\begin{aligned}
 & \text{bond broken} - \text{bond formed} = -90 \\
 & 1660 = x - 90 \\
 & \quad \quad \quad x = 1660 + 90 \\
 & \quad \quad \quad x = 1750 \text{ kJ}
 \end{aligned}$$
- Calculate the energy of one N-H bond.

$$\begin{aligned}
 & 6 \times (\text{N-H}) = 1750 \\
 & \text{N-H} = \frac{1750}{6} \quad \quad \quad 292 \text{ kJ/mol}
 \end{aligned}$$

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Roman Wölfel

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Table of Contents 0620 61 M J 14 Mark Scheme

1. Understanding the eBook 0620 61 M J 14 Mark Scheme
 - The Rise of Digital Reading 0620 61 M J 14 Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying 0620 61 M J 14 Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 0620 61 M J 14 Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from 0620 61 M J 14 Mark Scheme
 - Personalized Recommendations
 - 0620 61 M J 14 Mark Scheme User Reviews and Ratings
 - 0620 61 M J 14 Mark Scheme and Bestseller Lists
5. Accessing 0620 61 M J 14 Mark Scheme Free and Paid eBooks

- 0620 61 M J 14 Mark Scheme Public Domain eBooks
 - 0620 61 M J 14 Mark Scheme eBook Subscription Services
 - 0620 61 M J 14 Mark Scheme Budget-Friendly Options
6. Navigating 0620 61 M J 14 Mark Scheme eBook Formats
- ePub, PDF, MOBI, and More
 - 0620 61 M J 14 Mark Scheme Compatibility with Devices
 - 0620 61 M J 14 Mark Scheme Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of 0620 61 M J 14 Mark Scheme
 - Highlighting and Note-Taking 0620 61 M J 14 Mark Scheme
 - Interactive Elements 0620 61 M J 14 Mark Scheme
8. Staying Engaged with 0620 61 M J 14 Mark Scheme
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 0620 61 M J 14 Mark Scheme
9. Balancing eBooks and Physical Books 0620 61 M J 14 Mark Scheme
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection 0620 61 M J 14 Mark Scheme
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 0620 61 M J 14 Mark Scheme
- Setting Reading Goals 0620 61 M J 14 Mark Scheme
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 0620 61 M J 14 Mark Scheme
- Fact-Checking eBook Content of 0620 61 M J 14 Mark Scheme
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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