

Chemistry

Page 5 of 64

1A

FUNDAMENTAL CONCEPTS

5

2011

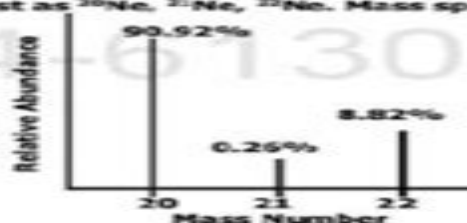
- Q.1 In mass spectrometer, detector or collector measures the:
 A) Masses of isotopes
 B) Percentages of isotopes
 C) Relative abundances of isotopes
 D) Mass numbers of isotopes
- Q.2 How many ^{35}Cl (chlorine) atoms are in two moles of chlorine?
 A) $2 \times 6.02 \times 10^{23}$ atoms
 B) $35.5 \times 6.02 \times 10^{23}$ atoms
 C) 2×16^{23} atoms
 D) $2 \times 6.02 \times 10^{23}$ atoms

2012

- Q.3 An organic compound has empirical formula $\text{C}_3\text{H}_5\text{O}$, if molar mass of compound is $110.15 \text{ g mol}^{-1}$. The molecular formula of this organic compound is (A, of C=12, H=1.008 and O=16)
 A) $\text{C}_6\text{H}_{10}\text{O}_2$
 B) $\text{C}_6\text{H}_{10}\text{O}$
 C) $\text{C}_3\text{H}_5\text{O}$
 D) $\text{C}_6\text{H}_5\text{O}_2$
- Q.4 When 8 grams (4 moles) of H_2 react with 2 moles of O_2 , how many moles of water will be formed?
 A) Five
 B) Four
 C) Six
 D) Three

2013

- Q.5 Hydrogen burns in chlorine to produce hydrogen chloride. The ratio of masses of reactants in chemical reaction is:
 $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 A) 1:35.5
 B) 2:35.5
 C) 1:71
 D) 2:70
- Q.6 A sample of Neon is found to exist as ^{20}Ne , ^{21}Ne , ^{22}Ne . Mass spectrum of 'Ne' is as follow:



What is the relative atomic mass (A, value) of Neon?

- A) 20.18
 B) 20.28
 C) 20.10
 D) 20.22

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Table of Contents **06biology Past Papers**

1. Understanding the eBook **06biology Past Papers**
 - The Rise of Digital Reading **06biology Past Papers**
 - Advantages of eBooks Over Traditional Books
2. Identifying **06biology Past Papers**
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an **06biology Past Papers**
 - User-Friendly Interface
4. Exploring eBook Recommendations from **06biology Past Papers**
 - Personalized Recommendations
 - **06biology Past Papers** User Reviews and Ratings
 - **06biology Past Papers** and Bestseller Lists

5. Accessing 06biology Past Papers Free and Paid eBooks
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 - 06biology Past Papers eBook Subscription Services
 - 06biology Past Papers Budget-Friendly Options
6. Navigating 06biology Past Papers eBook Formats
 - ePub, PDF, MOBI, and More
 - 06biology Past Papers Compatibility with Devices
 - 06biology Past Papers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 06biology Past Papers
 - Highlighting and Note-Taking 06biology Past Papers
 - Interactive Elements 06biology Past Papers
8. Staying Engaged with 06biology Past Papers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 06biology Past Papers
9. Balancing eBooks and Physical Books 06biology Past Papers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 06biology Past Papers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 06biology Past Papers
 - Setting Reading Goals 06biology Past Papers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 06biology Past Papers
 - Fact-Checking eBook Content of 06biology Past Papers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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