

Basic Chemistry Study Guide 2

1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's δ^+ region is attracted to a δ^- region in another molecule.
3. Form between water molecules (Where are the δ^+ and δ^- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

2. Solutions, solvents, and solutes

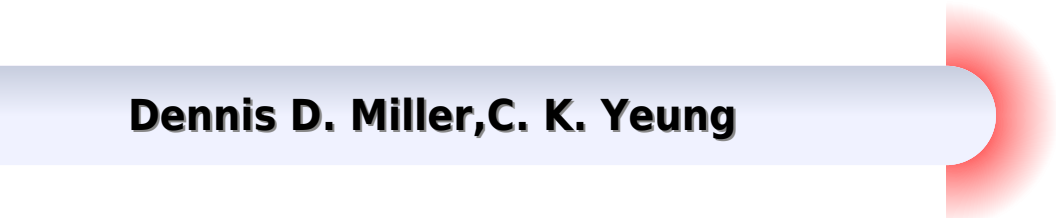
1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions
4. Define the terms **solution, solvent, and solute**. **Solvent: a substance capable of dissolving other polar molecules and ionic bonds, solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion, cation, and anion**. **ion: atom with a charge, cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water, hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar hydrophobic, polar hydrophilic**

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Dennis D. Miller, C. K. Yeung



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Table of Contents 100 Question Study Guide For Chemistry

1. Understanding the eBook 100 Question Study Guide For Chemistry
 - The Rise of Digital Reading 100 Question Study Guide For Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying 100 Question Study Guide For Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 100 Question Study Guide For Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from 100 Question Study Guide For Chemistry
 - Personalized Recommendations
 - 100 Question Study Guide For Chemistry User Reviews and Ratings
 - 100 Question Study Guide For Chemistry and Bestseller Lists
5. Accessing 100 Question Study Guide For Chemistry Free and Paid eBooks
 - 100 Question Study Guide For Chemistry Public Domain eBooks
 - 100 Question Study Guide For Chemistry eBook Subscription Services
 - 100 Question Study Guide For Chemistry Budget-Friendly Options

6. Navigating 100 Question Study Guide For Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - 100 Question Study Guide For Chemistry Compatibility with Devices
 - 100 Question Study Guide For Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 100 Question Study Guide For Chemistry
 - Highlighting and Note-Taking 100 Question Study Guide For Chemistry
 - Interactive Elements 100 Question Study Guide For Chemistry
8. Staying Engaged with 100 Question Study Guide For Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 100 Question Study Guide For Chemistry
9. Balancing eBooks and Physical Books 100 Question Study Guide For Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 100 Question Study Guide For Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine 100 Question Study Guide For Chemistry
 - Setting Reading Goals 100 Question Study Guide For Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of 100 Question Study Guide For Chemistry
 - Fact-Checking eBook Content of 100 Question Study Guide For Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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