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(a) The symbol equation for the reaction can be represented as shown in Fig. 3.1.

Fig. 3.1

Table 3.1 shows some bond energies.

Bond	H-H	N-N
Bond energy in kJ/mol	435	193

Use the bond energies in Table 3.1 and $\Delta H = -91 \text{ kJ/mol}$ to calculate the bond energy of an N-H bond in kJ/mol.

Use the following steps.

- Calculate the energy needed to break bonds in the reactants.

$$(4 \times 435) + 1(193)$$

$$4 \times 435 + 1(193) = 2150 \text{ kJ/mol}$$
- Calculate the energy released when bonds form in the products.

$$\text{Bond broken} - \text{Bond formed} = -91$$

$$2150 - x = -91$$

$$x = 2150 + 91$$

$$x = 2241 \text{ kJ/mol}$$
- Calculate the energy of per mole bond.

$$6 \times \text{N-H} = 2241$$

$$\text{N-H} = \frac{2241}{6} = 373.5 \text{ kJ/mol}$$

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