NATURAL SCIENCES

Gr. 9 - LESSON PLAN - MATTER & MATERIALS

	TOPIC: COMPOU	INDS	
Sub-topic: Names of Compounds		Duration: 1 Hour	
	CONTENT & CONCEPTS	(CAPS p63)	
KEY CONCEPTS			
1. Chemical	reaction		
Chemical	formula		
Names of	compounds according to their elements.		
Specific Aims:	Specific Aim 1: 'Doing Science'		X
Specific Airis.			
	Specific Aim 2: 'Knowing the subject content and making connections'		X
	Specific Aim 3: 'Understanding the uses	of Science'	

LESSON OBJECTIVES:

- Learners will be able to write the names of compounds as indicated on page 63 of the CAPS Document.
- 2. Learners will be able to make models of different compounds

RESOURCES REQUIRED:

- Periodic Table Poster
- Siyavula Sasol Inzalo (Learner book) page 163

TEACHING & LEARNING ACTIVITY: Introduction Briefly review and revise concepts dealt with in Grade 8, focusing on compounds

As each element has a unique name for example:

- Hydrogen
- Lithium

Remember the first 20 elements learnt in grade 8. How did you memorize them? Do you remember Mnemonics...

Hit - H; He - He; Lies - Li; Because - Be; Boys - B; Can - C; Not - N; Operate - O; Fireplaces - F;

New - No: Nation - No: Might - Mg; Also - Al; Sign - Si; Peace - P; Security - S; Clause - Cl

A - Ar; King - K; Can - Ca LEARNERS MAY RECITE THIS MINEMONIC AS A WAY OF MEMORIZING THE FIRST 20 ELEMENTS.

The teacher provides the following explanation to the learners

Each compound has a unique name. How is a compound formed?

- A compound is formed when TWO or more different elements combine chemically.
- Each atom in a compound are held by chemical bonds
- Compounds are formed during chemical reaction (when elements/ componds combine to form new substances)
- Each compound can be written as a **formula**, for example.

 H-O: NaCl: CuSO₄

What is a formula?

- This is a notation that shows which elements the compound is made up of.
- It indicates the total number of atoms in one unit of the compound
- It indicates the ratio in which atoms of different elements are bonded to each other.
 It gives the total number of atoms in one unit of a compound.

For example, the formula for Copper Sulphate is CuSO.

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Xiaolong Qi

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