

Science

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(Chapter – 1) (Chemical Reactions and Equations) (Exercises) (Class 10)

Question 1:

Which of the statements about the reaction below are incorrect?



- (a) Lead is getting reduced.
- (b) Carbon dioxide is getting oxidised.
- (c) Carbon is getting oxidised.
- (d) Lead oxide is getting reduced.

(i) (a) and (b)

(ii) (a) and (c)

(iii) (a), (b) and (c)

(iv) all

Answer 1:

(i) (a) and (b)

Question 2:

The above reaction is an example of a



- (a) combination reaction.
- (c) decomposition reaction.

- (b) double displacement reaction.
- (d) displacement reaction.

Answer 2:

(d) displacement reaction.

Question 3:

What happens when dilute hydrochloric acid is added to iron fillings? Tick the correct answer.

- (a) Hydrogen gas and iron chloride are produced.
- (c) No reaction takes place.

- (b) Chlorine gas and iron hydroxide are produced.
- (d) Iron salt and water are produced.

Answer 3:

(a) Hydrogen gas and iron chloride are produced.

Question 4:

What is a balanced chemical equation? Why should chemical equations be balanced?

Answer 4:

A reaction which has an equal number of atoms of all the elements on both sides of the chemical equation is called a balanced chemical equation.

The law of conservation of mass states that mass can neither be created nor destroyed. Hence, in a chemical reaction, the total mass of reactants should be equal to the total mass of the products. It means that the total number of atoms of each element should be equal on both sides of a chemical equation. Hence, it is for this reason that chemical equations should be balanced.

Question 5:

Translate the following statements into chemical equations and then balance them.

- (a) Hydrogen gas combines with nitrogen to form ammonia.
- (b) Hydrogen sulphide gas burns in air to give water and sulphur dioxide.
- (c) Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate.
- (d) Potassium metal reacts with water to give potassium hydroxide and hydrogen gas.

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Mysore (India : State)



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1. Understanding the eBook Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
 - The Rise of Digital Reading Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
 - Personalized Recommendations
 - Answers Of 12 Th Science 18 03 14 Dated Maths Papaer User Reviews and Ratings
 - Answers Of 12 Th Science 18 03 14 Dated Maths Papaer and Bestseller Lists
5. Accessing Answers Of 12 Th Science 18 03 14 Dated Maths Papaer Free and Paid eBooks
 - Answers Of 12 Th Science 18 03 14 Dated Maths Papaer Public Domain eBooks
 - Answers Of 12 Th Science 18 03 14 Dated Maths Papaer eBook Subscription Services
 - Answers Of 12 Th Science 18 03 14 Dated Maths Papaer Budget-Friendly Options
6. Navigating Answers Of 12 Th Science 18 03 14 Dated Maths Papaer eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Of 12 Th Science 18 03 14 Dated Maths Papaer Compatibility with Devices
 - Answers Of 12 Th Science 18 03 14 Dated Maths Papaer Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
 - Highlighting and Note-Taking Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
 - Interactive Elements Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
8. Staying Engaged with Answers Of 12 Th Science 18 03 14 Dated Maths Papaer

- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
9. Balancing eBooks and Physical Books Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
- Setting Reading Goals Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
- Fact-Checking eBook Content of Answers Of 12 Th Science 18 03 14 Dated Maths Papaer
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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