



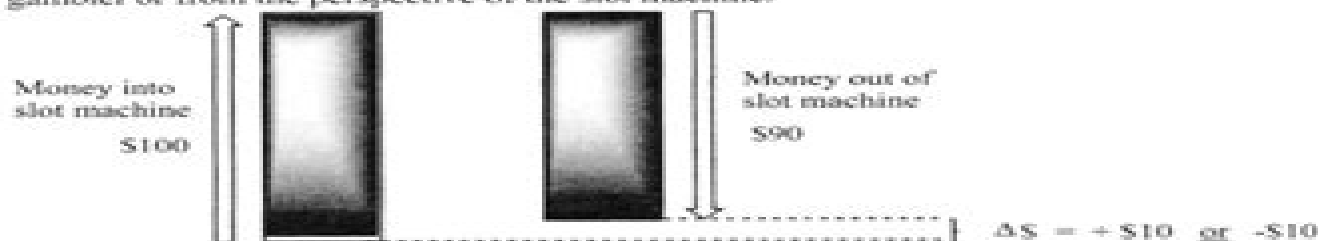
## ChemActivity 6

### Part A: Energy Diagrams

(What is the sign of  $\Delta H_{\text{reaction}}$  for an exothermic reaction?)

#### Model 1A: The Slot Machine

Imagine a gambler sitting at a slot machine. Over the course of an hour the gambler puts \$100 into the machine and the machine pays out \$90. We know a net change ( $\Delta$ ) of ten dollars occurred, but what is the sign of this change in money ( $\Delta S$ ). To determine if  $\Delta S = +10$  or  $-10$  we need to decide if we calculate  $\Delta S$  from the perspective of the gambler or from the perspective of the slot machine.



To translate this analogy into chemistry remember that:

- **Money = energy.**
- **The slot machine = the reaction.**
- **The gambler = the experimenter (you).**
- **All changes ( $\Delta$ ) will be calculated from the perspective of the reaction (a.k.a. the slot machine).**

#### Critical Thinking Questions

1. Using the convention listed above for calculating changes ( $\Delta$ ), what is the sign of  $\Delta S$  in the example in Model 1A?
2. Based on this same convention, what sign must be attached to the \$100 that the gambler put into the slot machine in Model 1A? The sign of the \$90 that comes out is the opposite. **Add the appropriate sign to each value in Model 1A.**
3. On a different day the gambler puts in a total of \$20 and the slot machine pays out a total of \$100. What is the sign and value of  $\Delta S$  for this day?



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## **Table of Contents Answers Chemactivity 6**

1. Understanding the eBook Answers Chemactivity 6
  - The Rise of Digital Reading Answers Chemactivity 6
  - Advantages of eBooks Over Traditional Books
2. Identifying Answers Chemactivity 6
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Answers Chemactivity 6
  - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Chemactivity 6
  - Personalized Recommendations
  - Answers Chemactivity 6 User Reviews and Ratings
  - Answers Chemactivity 6 and Bestseller Lists
5. Accessing Answers Chemactivity 6 Free and Paid eBooks
  - Answers Chemactivity 6 Public Domain eBooks
  - Answers Chemactivity 6 eBook Subscription Services
  - Answers Chemactivity 6 Budget-Friendly Options
6. Navigating Answers Chemactivity 6 eBook Formats

- ePub, PDF, MOBI, and More
- Answers Chemactivity 6 Compatibility with Devices
- Answers Chemactivity 6 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Answers Chemactivity 6
  - Highlighting and Note-Taking Answers Chemactivity 6
  - Interactive Elements Answers Chemactivity 6
- 8. Staying Engaged with Answers Chemactivity 6
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Answers Chemactivity 6
- 9. Balancing eBooks and Physical Books Answers Chemactivity 6
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Answers Chemactivity 6
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Chemactivity 6
  - Setting Reading Goals Answers Chemactivity 6
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Chemactivity 6
  - Fact-Checking eBook Content of Answers Chemactivity 6
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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