

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Not all orbitals have the same amount of energy. For example, orbitals closer to the nucleus have more energy than orbitals further from the nucleus. The following orbitals have different levels of energy:



For each pair, write whether one orbital is higher in energy than the other orbital (higher or lower).



Now list 8 orbitals which appear in the set of pairs. Write down their approximate energy. For example, the 3d orbitals have higher energy than 4s orbitals but 4f orbitals have a higher energy than 3d orbitals.

When electrons occupy orbitals, they try to have the lowest amount of energy possible. (This is called the **lowest-energy principle**.) An electron will never be added to orbitals that are higher in energy than the orbitals it is already in. Additionally, two electrons will not occupy the same orbital. (This is called the **Pauli Exclusion Principle**.) When two electrons occupy the same orbital they must spin in opposite directions—one clockwise and the other anticlockwise.

Critical Thinking Questions

- a) How many electrons would it take to fill orbitals from 1s up to 4s orbitals? **(Orbitals from 1s up to 4s orbitals are filled with a maximum of 20 electrons.)**
b) What is the first element that has enough electrons to fill up to the 4s orbitals? **(This is Potassium and the first element that has 4s orbitals filled is Calcium.)**
- a) How many electrons would it take to fill orbitals from 1s up to 4p orbitals? **(Orbitals from 1s up to 4p orbitals are filled with a maximum of 18 electrons.)**
b) What is the first element that has enough electrons to fill up to the 4p orbitals? **(This is Nitrogen and the first element that has 4p orbitals filled is Argon.)**

Information: Hund's Rule

Electrons prefer to be "paired" or "coupled". Orbitals closer to the nucleus have a higher potential energy than orbitals further from the nucleus. For example, 3s orbitals have a higher potential energy than 3p orbitals.



If an orbital is empty then an electron's orbital energy will go down to the orbital's energy. The higher the energy is, the lower the orbital energy. The electron has a higher energy than the orbital energy—this is the 3p orbital's energy.



Hund's rule tells us that orbitals closer to the nucleus have a lower energy than orbitals further from the nucleus. Orbitals closer to the nucleus have a lower energy than orbitals further from the nucleus. Orbitals closer to the nucleus have a lower energy than orbitals further from the nucleus. Orbitals closer to the nucleus have a lower energy than orbitals further from the nucleus.

The following are the electron orbital diagrams for the first elements, hydrogen and helium. Write the electron orbital diagram for the first element, hydrogen and helium. The 1s orbital is the first orbital to be filled.



Critical Thinking Questions

- How many "paired" electrons are in a nitrogen atom? **(1)**
- Why does carbon's electron distribution go from 1s to 2s to 2p? Why not 1s to 2s to 3s? **(The 3s orbital is already full, so it is not available for the 2s orbital to fill.)**
- Why do elements like nitrogen and oxygen have 2s orbitals? **(The 2s orbital is already full, so it is not available for the 2s orbital to fill.)**
- Why do elements like nitrogen and oxygen have 2p orbitals? **(The 2p orbital is already full, so it is not available for the 2p orbital to fill.)**

Answers For Chemquest

RJ Alexander



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Chemistry Buchanan, 2000-12-01

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