

Name: _____

Chemistry Practice: Balancing Chemical Equations

1. Balance the equations below. Remember this means only using coefficients. All formulas are written correctly.

- a.) $2 \text{NaHCO}_3 \longrightarrow \text{Na}_2\text{CO}_3 + \text{H}_2\text{O} + \text{CO}_2$
- b.) $\text{Zn} + \text{S} \longrightarrow \text{ZnS}$
- c.) $2 \text{H}_2 + \text{O}_2 \longrightarrow 2 \text{H}_2\text{O}$
- d.) $2 \text{FeO} + \text{C} \longrightarrow 2 \text{Fe} + \text{CO}_2$
- e.) $\text{Cl}_2 + 2 \text{KI} \longrightarrow 2 \text{KCl} + \text{I}_2$
- f.) $\text{B}_2\text{O}_3 + 3 \text{Mg} \longrightarrow 3 \text{MgO} + 2 \text{B}$
- g.) $\text{Ca} + 2 \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2 + 2 \text{H}_2$
- h.) $4 \text{H}_2\text{O} + \text{Na}_2\text{B}_4\text{O}_7 + \text{H}_2\text{SO}_4 \longrightarrow 4 \text{H}_3\text{BO}_3 + \text{Na}_2\text{SO}_4$
- i.) $\text{CaC}_2 + 2 \text{H}_2\text{O} \longrightarrow \text{Ca(OH)}_2 + \text{C}_2\text{H}_2$
- j.) $\text{NH}_4\text{NO}_3 \longrightarrow \text{N}_2\text{O} + 2 \text{H}_2\text{O}$
- k.) $\text{H}_2\text{S} + \text{O}_2 \longrightarrow \text{SO}_2 + \text{H}_2\text{O}$
- l.) $\text{H}_2\text{SO}_4 + 2 \text{NaCl} \longrightarrow 2 \text{HCl} + \text{Na}_2\text{SO}_4$
- m.) $2 \text{NaCl} + 2 \text{H}_2\text{O} \longrightarrow 2 \text{NaOH} + \text{Cl}_2 + \text{H}_2$
- n.) $3 \text{BaO} + 2 \text{Al} \longrightarrow 3 \text{Ba} + \text{Al}_2\text{O}_3$
- o.) $3 \text{O}_2 \longrightarrow 2 \text{O}_3$
- p.) $\text{SiO}_2 + 2 \text{C} \longrightarrow 2 \text{CO} + \text{Si}$
- q.) $2 \text{PbS} + 3 \text{O}_2 \longrightarrow 2 \text{PbO} + 2 \text{SO}_2$
- r.) $2 \text{HgO} \longrightarrow 2 \text{Hg} + \text{O}_2$
- s.) $\text{C}_{12}\text{H}_{22}\text{O}_{11} + 12 \text{O}_2 \longrightarrow 12 \text{CO}_2 + 11 \text{H}_2\text{O}$
- t.) $\text{C}_3\text{H}_8 + 5 \text{O}_2 \longrightarrow 3 \text{CO}_2 + 4 \text{H}_2\text{O}$
- u.) $2 \text{FeCl}_3 + 3 (\text{NH}_4)_2\text{S} \longrightarrow 6 \text{NH}_4\text{Cl} + \text{Fe}_2\text{S}_3$
- v.) $\text{P}_2\text{O}_5 + 6 \text{NaOH} \longrightarrow 2 \text{Na}_3\text{PO}_4 + 3 \text{H}_2\text{O}$
- w.) $3 \text{BaCl}_2 + \text{Al}_2(\text{SO}_4)_3 \longrightarrow 3 \text{BaSO}_4 + 2 \text{AlCl}_3$
- x.) $4 \text{NH}_3 + 5 \text{O}_2 \longrightarrow 6 \text{H}_2\text{O} + 4 \text{NO}$
- y.) $2 \text{Na}_2\text{O}_2 + 2 \text{H}_2\text{O} \longrightarrow 4 \text{NaOH} + \text{O}_2$
- z.) $\text{Sn(ClO}_3)_4 \longrightarrow \text{SnCl}_4 + 6 \text{O}_2$

Text Problems to be Completed on the Other Side of this Worksheet:

*P. 173 #45,49,53,54,56,57,58,63, 65,66,67,71

*(Although the directions may say to write an "unbalanced" equation, write a balanced equation for each.)

10th Grade Chemistry Starr Practice Test

Reza Nazari



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Table of Contents 10th Grade Chemistry Starr Practice Test

1. Understanding the eBook 10th Grade Chemistry Starr Practice Test
 - The Rise of Digital Reading 10th Grade Chemistry Starr Practice Test
 - Advantages of eBooks Over Traditional Books
2. Identifying 10th Grade Chemistry Starr Practice Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 10th Grade Chemistry Starr Practice Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from 10th Grade Chemistry Starr Practice Test

- Personalized Recommendations
- 10th Grade Chemistry Starr Practice Test User Reviews and Ratings
- 10th Grade Chemistry Starr Practice Test and Bestseller Lists
- 5. Accessing 10th Grade Chemistry Starr Practice Test Free and Paid eBooks
 - 10th Grade Chemistry Starr Practice Test Public Domain eBooks
 - 10th Grade Chemistry Starr Practice Test eBook Subscription Services
 - 10th Grade Chemistry Starr Practice Test Budget-Friendly Options
- 6. Navigating 10th Grade Chemistry Starr Practice Test eBook Formats
 - ePub, PDF, MOBI, and More
 - 10th Grade Chemistry Starr Practice Test Compatibility with Devices
 - 10th Grade Chemistry Starr Practice Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 10th Grade Chemistry Starr Practice Test
 - Highlighting and Note-Taking 10th Grade Chemistry Starr Practice Test
 - Interactive Elements 10th Grade Chemistry Starr Practice Test
- 8. Staying Engaged with 10th Grade Chemistry Starr Practice Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 10th Grade Chemistry Starr Practice Test
- 9. Balancing eBooks and Physical Books 10th Grade Chemistry Starr Practice Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 10th Grade Chemistry Starr Practice Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine 10th Grade Chemistry Starr Practice Test
 - Setting Reading Goals 10th Grade Chemistry Starr Practice Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of 10th Grade Chemistry Starr Practice Test

- Fact-Checking eBook Content of 10th Grade Chemistry Starr Practice Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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