

Appendix G

Answers to Odd-Numbered Problems

Chapter 1

1.1 (a) -0.1038 C , (b) -0.19865 C , (c) -3.941 C ,
(d) -26.08 C

1.3 (a) $3t + 1\text{ C}$, (b) $t^2 + 5t\text{ mC}$,
(c) $2\sin(10t + \pi/6) + 1\text{ }\mu\text{C}$,
(d) $-e^{-30t}[0.16\cos 40t + 0.12\sin 40t]\text{ C}$

1.5 25 C

1.7 $i = \begin{cases} 25\text{ A}, & 0 < t < 2 \\ -25\text{ A}, & 2 < t < 6 \\ 25\text{ A}, & 6 < t < 8 \end{cases}$

See the sketch in Fig. G.1.

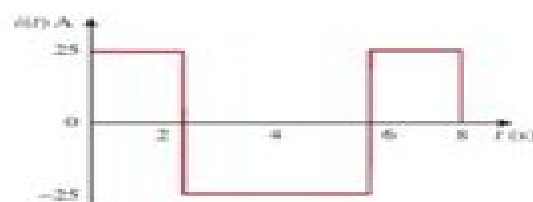


Figure G.1
For Prob. 1.7.

1.9 (a) 10 C , (b) 22.5 C , (c) 30 C

1.11 3.672 kC , 4.406 kJ

1.13 164.5 mW , 78.34 mJ

1.15 (a) 1.297 C , (b) $-90e^{-40t}\text{ W}$, (c) -22.5 J

1.17 70 W

1.19 3 A

1.21 $2.696 \times 10^{23}\text{ electrons}$, $43,200\text{ C}$

1.23 51.35

1.25 21.6 cents

1.27 (a) 43.2 kC , (b) 475.2 kJ , (c) 1.188 cents

1.29 39.6 cents

1.31 $\$42.05$

1.33 6 C

1.35 2.333 MWh

1.37 29.84 kWh

1.39 24 cents

Chapter 2

2.1 This is a design problem with several answers.

2.3 184.3 mm

2.5 $n = 9$, $b = 15$, $f = 7$

2.7 (a) 6 branches and 5 nodes, and (b) 7 branches and 5 nodes.

2.9 14 A , -2 A , 10 A

2.11 6 V , 3 V

2.13 12 A , -10 A , 5 A , -2 A

2.15 10 V , -2 A

2.17 2 V , -22 V , 10 V

2.19 2 A , 12 W , 8 W , -40 W , 20 W

2.21 4.167 W

2.23 2 V , 1.92 W

2.25 0.1 A , 2 kV , 0.2 kW

2.27 6.4 V

2.29 $1.625\text{ }\Omega$

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S Nieto



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for the nitration 4 01 ml of nitric acid 4 00 ml of sulfuric acid and 4 56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4 56 ml of bromobenzene was used the amount of product predicted to be recovered was 8 7533 g

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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over 5 minutes slowly pipet this bromobenzene into the icy cold nitration mixture hno 3 h 2 so 4 in the 125 ml flask and stir continue to stir the reaction mixture in the ice bath for 30 minutes do not heat in a new 250 ml erlenmeyer

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on the right the arenium cation has a nitro set with incredibly small 0.6 angstrom bond distances this will result in an optimization failure as the atoms are even close for the programme to accomodate their electrons demonstration of creating the p arenium cation by to nitration of bromobenzene

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