

Stoichiometry Practice Problems

Name: _____

Date: _____ Period: _____

Reminder...Quiz next class on the following types of problems. Make sure you understand: 1. Molar mass 2. Mole Ratio 3. Avogadro's number.

Moles to Moles

1. Given the *unbalanced* equation: $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$

a) If 4.50 moles of C_2H_6 are used, how many moles of oxygen are required?

b) If 0.3 moles of H_2O were produced, how many moles of carbon dioxide would also have been produced?

c) If you wanted to produce 5 moles of carbon dioxide, how many moles of oxygen would you needed to have started with?

Grams to Moles

2. Given the *unbalanced* equation below:



a) How many grams of H_2SO_4 are needed to react with 0.75 moles of NaOH?

b) How many grams of Na_2SO_4 would be formed if 0.75 moles of NaOH reacted?

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Silberberg



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Prentice Hall Chemistry, 2000 **A Stoichiometry Unit** David Callaghan, 2004 *Cehmistry Textbook for College and University USA* Ibrahim Sikder, 2023-06-04 Cehmistry Textbook USA **Jacaranda Chemistry 2 VCE Units 3 and 4, 3e learnON and Print** Robert Stokes, Angela Stubbs, Neale Taylor, Jason Bourke, Ben Williams, Maida Derbogolian, 2023-12-26 Developed by expert Victorian teachers for VCE students The NEW Jacaranda Chemistry VCE series continues to deliver curriculum aligned material that caters to students of all abilities Our expert author team of practising teachers and assessors ensures 100% coverage of the new VCE Chemistry Study Design 2023 2027 **Chemistry Class 11 - [Bihar & JAC]** Dr. S. C. Rastogi, ,Er. Meera Goyal, 2022-07-20 Syllabus Unit I Some Basic Concepts of Chemistry Unit II Structure of Atom Unit III Classification of Elements and Periodicity in Properties Unit IV Chemical Bonding and Molecular Structure Unit V States of Matter Gases and Liquids Unit VI Chemical Thermodynamics Unit VII Equilibrium Unit VIII Redox Reactions Unit IX Hydrogen Unit X s Block Elements Alkali and Alkaline earth metals Group 1 and Group 2 Elements Unit XI Some p Block Elements General Introduction to p Block Elements Unit XII Organic Chemistry Some Basic Principles and Techniques Unit XIII Hydrocarbons Classification of Hydrocarbons Unit XIV Environmental Chemistry Content 1 Some Basic Concepts of Chemistry 2 Structure of Atom 3 Classification of Elements and Periodicity in Properties 4 Chemical Bonding and Molecular Structure 5 States of Matter 6 Thermodynamics 7 Equilibrium 8 Redox Reactions 9 Hydrogen 10 s Block Elements 11 p Block Elements 12 Organic Chemistry Some Basic Principles and Techniques 13 Hydrocarbons 14 Environmental Chemistry I Appendix II Log antilog Table *CliffsNotes Chemistry Practice Pack* Charles Henrickson, 2010-02-08 About the Contents Pretest Helps you pinpoint where you need the most help Topic Area Reviews Measurement and Units of Measurement Matter Elements Compounds and Mixtures Atoms I The Basics Formulas and Names of Ionic Compounds Acids and Bases The Mole Elements and Compounds Percent Composition and Empirical and Molecular Formulas Chemical Reactions and Chemical Equations Calculations Using Balanced Equations Atoms II Atomic Structure and Periodic Properties Chemical Bonding The Formation of Compounds Gases and the Gas Laws The Forces between Molecules Solids and Liquids Solutions and Solution Composition Acids Bases and Neutralization Glossary Customized Full Length Exam Covers all subject areas Pretest that pinpoints what you need to study most Clear concise reviews of every topic Targeted example problems in every chapter with solutions and explanations Customized full length exam that adapts to your skill level **NCERT Chemistry Class 11 - [CBSE Board]** Dr. S. C. Rastogi, ,Er. Meera Goyal, 2022-10-11 Syllabus Unit I Some Basic Concepts of Chemistry Unit II Structure of Atom Unit III Classification of Elements and Periodicity in Properties Unit IV Chemical Bonding and Molecular Structure Unit V States of Matter Gases and Liquids Unit VI Chemical Thermodynamics Unit VII Equilibrium Unit VIII Redox Reactions Unit IX Hydrogen Unit X s Block Elements Alkali and Alkaline earth metals Group 1 and Group 2 Elements Unit XI Some p Block Elements General Introduction to p Block Elements Unit XII Organic Chemistry Some Basic

Principles and Techniques Unit XIII Hydrocarbons Classification of Hydrocarbons Unit XIV Environmental Chemistry Content
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Chemical Bonding and Molecular Structure 5 States of Matter 6 Thermodynamics 7 Equilibrium 8 Redox Reactions 9
Hydrogen 10 s Block Elements 11 p Block Elements 12 Organic Chemistry Some Basic Principles and Techniques 13
Hydrocarbons 14 Environmental Chemistry I Appendix II Log antilog Table **Ebook: Chemistry: The Molecular Nature
of Matter and Change** Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change Essentials
of Chemical Reaction Engineering H. Scott Fogler, 2017-10-26 Today's Definitive Undergraduate Level Introduction to
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has been the 1 selling text for courses in chemical reaction engineering worldwide Now in Essentials of Chemical Reaction
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Basic Concepts of Chemistry Leo J. Malone, Theodore

Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book. It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features. Each section focuses on the development and assessment of one or two specific objectives. Within each section a specific objective is included, an anticipatory set to orient the reader, content discussion from established authors, and guided practice problems for relevant objectives. These features are followed by a set of independent practice problems. The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics. Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance.

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