

Chapter 11 – Thermochemistry – Heat and Chemical Change

Chapter 11: 1 – 35, 57, 60, 61, 71

Section 11.1 – The Flow of Energy - Heat

Practice Problems

1. When 435 J of heat is added to 3.4 g of olive oil at 21°C, the temperature increases to 85°C. What is the specific heat of olive oil?

Knowns: $q = 435 \text{ J}$; $m_{\text{olive oil}} = 3.4 \text{ g}$; $\Delta T = (85^\circ\text{C} - 21^\circ\text{C})$

Use $C = q / m \cdot \Delta T$

$$C = \frac{435 \text{ J}}{3.4 \text{ g} \cdot 64^\circ\text{C}} = 2.0 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$$

2. A 1.55-g piece of stainless steel absorbs 141 J of heat when its temperature increases by 178°C. What is the specific heat of the stainless steel?

Knowns: $m_{\text{stainless steel}} = 1.55 \text{ g}$; $q = 141 \text{ J}$; $\Delta T = 178^\circ\text{C}$

Use $C = q / m \cdot \Delta T$

$$C = \frac{141 \text{ J}}{1.55 \text{ g} \cdot 178^\circ\text{C}} = 0.511 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}$$

3. How much heat is required to raise the temperature of 250.0 g of mercury 52°C?

Knowns: $m_{\text{Hg}} = 250.0 \text{ g}$; $\Delta T = 52^\circ\text{C}$; $C_{\text{Hg}} = 0.14 \text{ J/g} \cdot ^\circ\text{C}$

Use $C = q / m \cdot \Delta T$; rearrange to solve for q ; $q = C \cdot m \cdot \Delta T$

$$q = \left(0.14 \frac{\text{J}}{\text{g} \cdot ^\circ\text{C}}\right) (250.0 \text{ g})(52^\circ\text{C}) = 1,820 \text{ J}$$

Section Review 11.1

4. Define *energy* and explain how *energy* and *heat* are related.

Energy is the capacity for doing work or supplying heat. Heat is energy that is transferred between objects, when their temperatures are different.

5. Explain the difference between *heat capacity* and *specific heat*.

The specific heat of a substance is independent of its mass, whereas the heat capacity of an object is proportional to its mass.

6. Will the specific heat of 50 g of a substance be the same as, or greater than, the specific heat of 10 g of the same substance? Explain.

The specific heat between the two substances will be the same. Specific heat is not dependent upon the mass, whereas *heat capacity* is.

7. On a sunny day, why does the concrete deck around an outdoor swimming pool become hot, while the water stays cool?

The water stays cooler because it has a higher heat capacity than concrete. The sun's heat raises the temperature of the concrete more than that of the water.

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JG Myers



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