

Multiple Choice Questions (Topic Wise)

Four possible answers to each statement are given below. Tick (✓) mark the correct/best one.

Kinetic Theory of Gases

The branch of physics which deals with the transformation of heat energy into other forms of energy is called

- (a) mechanics (b) thermodynamics (c) nuclear physics (d) particle physics

The behaviour of gases is well accounted by the kinetic theory based on

- (a) microscopic approach (b) macroscopic approach (c) both of these (d) none of these

"Pressure everywhere inside the vessel will be the same provided the gas is of uniform density" is the statement of

- (a) Newton (b) Boyle (c) Charles (d) Pascal

The expression for pressure exerted by the gas on container on any side is

- (a) $P = \frac{1}{2} \rho \langle v^2 \rangle$ (b) $P = \frac{3}{2} \rho \langle v^2 \rangle$ (c) $P = \frac{1}{3} \rho \langle v^2 \rangle$ (d) $P = \frac{1}{3} \rho \langle v \rangle$

The expression for pressure exerted by 'N' molecules is

- (a) $P = \frac{2}{3} \frac{N}{V} \langle \frac{1}{2} mv^2 \rangle$ (b) $P = \frac{3}{2} \frac{N}{V} \langle \frac{1}{2} mv^2 \rangle$ (c) $P = \frac{1}{2} \frac{N}{V} \langle \frac{1}{2} mv^2 \rangle$ (d) $P = \frac{N}{V} \langle \frac{1}{2} mv^2 \rangle$

The ideal gas law is

- (a) $PT = NVK$ (b) $P = NKT$ (c) $PV = nRT$ (d) $P = nRT$

The value of boltzman constant 'K' is

- (a) $1.38 \times 10^{-23} \text{ JK}^{-1}$ (b) $1.38 \times 10^{-25} \text{ JK}^{-1}$ (c) $1.38 \times 10^{-27} \text{ JK}^{-1}$ (d) $1.38 \times 10^{-19} \text{ JK}^{-1}$

Real gases obey gas laws only at

- (a) low pressure and high temperature (b) high pressure and low temperature
(c) low pressure and low temperature (d) high pressure and high temperature

The value of universal gas constant '(R)' is

- (a) $6.314 \text{ J mol}^{-1} \text{ K}^{-1}$ (b) $7.314 \text{ J mol}^{-1} \text{ K}^{-1}$ (c) $8.314 \text{ J mol}^{-1} \text{ K}^{-1}$ (d) $9.314 \text{ J mol}^{-1} \text{ K}^{-1}$

Absolute temperature of an ideal gas is directly proportional to

- (a) $\langle P.E \rangle$ (b) $\langle K.E \rangle$ (c) both of these (d) none of these

According to Charles law

- (a) $PT = \text{Constant}$ (b) $P \propto \frac{1}{T}$ (c) $V \propto n$ (d) $V \propto T$

The average translational K.E. of molecules in a gas at temperature 27°C is

- (a) $6.21 \times 10^{-21} \text{ J}$ (b) $6.21 \times 10^{-19} \text{ J}$ (c) $6.21 \times 10^{-5} \text{ J}$ (d) $6.21 \times 10^{-15} \text{ J}$

Internal Energy, Work and Heat

The value of Avogadro's number ' N_A ' is

- (a) 6.022×10^{21} (b) 6.022×10^{23} (c) 6.022×10^{25} (d) 6.022×10^{27}

In the study of thermodynamics, the gas which is usually considered as a working substance is

- (a) ideal gas (b) real gas (c) either of these (d) none of these

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