

Chm 113 Final Study Guide Questions & Answers 100% Correct.

What is the $[H^+]$ of a solution with a pH of 3.78? - ANSWER
 $pH = 3.78$
 $[H^+] = 10^{-3.78}$
 $= 1.66 \times 10^{-4}$

If HX represents the formula of a weak acid, what is the formula of the conjugate base of HX? - ANSWER
If HX weak acid then the formula of the conjugate base of X-

Let HX represent the formula of a weak acid. Write a balanced chemical equation for the reaction of this weak acid with the strong base NaOH? You may omit any spectator ions if you wish. - ANSWER
 $HX + NaOH \rightarrow NaX + H_2O$
For the equation you wrote above, what two terms are considered/assumed equal? - ANSWER
 $[HA] = [A^-]$

Regarding today's experiment, what is the weak acid we are going to titrate and which strong base is used to titrate this weak acid? - ANSWER
Weak Acid: Acetic Acid
Strong Base: NaOH

If a weak acid, HX, is titrated with a strong base such as NaOH, will the equivalence point in the titration occur below pH 7.0, at pH 7.0 or above pH 7.0? - ANSWER
the equivalence point in the titration will be somewhat above pH 7.0

A student is presented with a 0.65 M solution of an unknown weak acid. The student measures the pH of that solution and determines a value of 3.78 for the unknown solution. What, then, is the K_a of this weak acid? - ANSWER
 $K_a = \frac{(1.66 \times 10^{-4})^2}{(.65 - 1.66 \times 10^{-4})} = 4.2 \times 10^{-8}$

What is the acid dissociation constant (K_a)? - ANSWER
 $K_a = \frac{[H_3O^+][A^-]}{[HA]}$

What is the pH of a weak acid whose K_a is 6.03×10^{-6} , and the ratio of $[A^-]/[HA]$ is 3. - ANSWER
 $6.03 \times 10^{-6} = \frac{[9/3]}{3}$
 $pH = 5.69$

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Albert A Gayle



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