

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

1. Explain how a chemical equation is written.
2. Explain how to write a balanced equation.
3. Write the steps for writing a balanced chemical equation.

Vocabulary

- | | |
|----------------------|----------------------|
| 1. chemical equation | 4. coefficient |
| 2. stoichiometric | 5. balanced equation |
| 3. reactant | |

Part A Completion

Use the vocabulary words in each sentence, writing either the correct word form or the correct chemical equation. Write each answer on a separate sheet of paper and underline the answer.

1. A chemical equation that includes information about the _____ of the reactants, the products, and the states of matter is called a _____ equation. The word _____ refers to the relative amounts of the _____ involved in the reaction.
2. A chemical equation is _____ if the number of atoms of each element is the same on both sides of the equation.
3. A chemical equation is _____ if the number of molecules of each reactant is the same as the number of molecules of each product.
4. A chemical equation is _____ if the number of atoms of each element is the same on both sides of the equation.
5. A chemical equation is _____ if the number of molecules of each reactant is the same as the number of molecules of each product.

6. The chemical equation for the reaction of hydrogen gas with oxygen gas is _____.
7. The chemical equation for the reaction of hydrogen gas with oxygen gas is _____.
8. The chemical equation for the reaction of hydrogen gas with oxygen gas is _____.
9. The chemical equation for the reaction of hydrogen gas with oxygen gas is _____.
10. The chemical equation for the reaction of hydrogen gas with oxygen gas is _____.

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11 1 Describing Chemical Reactions Section Review

Answer Key

Linda Swisher, Kevin T. Patton



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