

Types of Chemical Reactions Worksheet #1

H/Chemistry

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

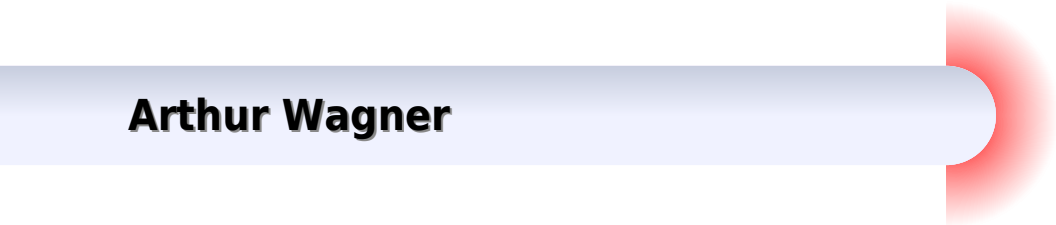
- 1) $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$ **Synthesis**
 - 2) $\text{CaI}_2 + \text{Cl}_2 \rightarrow \text{CaCl}_2 + \text{I}_2$ **Single Replacement**
 - 3) $3\text{KOH} + \text{AlCl}_3 \rightarrow \text{Al(OH)}_3 + 3\text{KCl}$ **Double Replacement**
 - 4) $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ **Synthesis**
 - *5) $\text{Ca(OH)}_2 + 2\text{HCl} \rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$ **Double Replacement**
 - 6) $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$ **Decomposition**
 - 7) $3\text{CuSO}_4 + 2\text{Al} \rightarrow \text{Al}_2(\text{SO}_4)_3 + 3\text{Cu}$ **Single Replacement**
 - 8) $\text{Na}_2\text{S} + 2\text{AgNO}_3 \rightarrow 2\text{NaNO}_3 + \text{Ag}_2\text{S}$ **Double Replacement**
 - 9) $3\text{H}_2 + \text{N}_2 \rightarrow 2\text{NH}_3$ **Synthesis**
 - *10) $2\text{Na} + 2\text{H}_2\text{O} \rightarrow 2\text{NaOH} + \text{H}_2$ **Single Replacement**
- *Remember: "H₂O" can also be thought of as "HOH"

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- 11) Double replacement: $\text{Na}_2\text{CrO}_4 + \text{PbCl}_2 \rightarrow 2\text{NaCl} + \text{PbCrO}_4$
- 12) Single replacement: $\text{Cl}_2 + 2\text{NaBr} \rightarrow 2\text{NaCl} + \text{Br}_2$
- 13) Decomposition: $\text{Mg(ClO}_3)_2 \rightarrow \text{MgCl}_2 + 3\text{O}_2$
- 14) Synthesis: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$
- 15) Double replacement: $3\text{Ca(OH)}_2 + 2\text{FeCl}_3 \rightarrow 2\text{Fe(OH)}_3 + 3\text{CaCl}_2$
- 16) Single replacement: $\text{Fe} + \text{Cu(NO}_3)_2 \rightarrow \text{Fe(NO}_3)_2 + \text{Cu}$ [Assume Fe²⁺]
- 17) Decomposition: $2\text{Hg}_2\text{O} \rightarrow 4\text{Hg} + \text{O}_2$
- 18) Synthesis: $\text{S} + \text{O}_2 \rightarrow \text{SO}_2$
- 19) Double replacement: $\text{AgNO}_3 + \text{KI} \rightarrow \text{AgI} + \text{KNO}_3$
- 20) Single replacement: $\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Ag} + \text{Cu(NO}_3)_2$ [Copper (II) is used here]

11describing Chemical Reactions Answers

Arthur Wagner



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Qualitative Theory of Parabolic Equations, Part 1 T. I. Zelenyak, M. P. Vishnevskii, M. M. Lavrentiev, 2011-09-06 No detailed description available for Qualitative Theory of Parabolic Equations Part 1 Conceptual Chemistry Volume I For Class XI S.K. Jain & Shailesh K. Jain, 1998 Conceptual Chemistry Volume I For Class XI **General Chemistry** Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value **Interactive Science For Inquiring Minds Volume B Theory Workbook**

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traditional field of control theory in scientific community **Pre-GED Science** Arthur Wagner, 1999-07-20 Comprehensive preparation for the Science of the GED Covers thoroughly the areas of biology chemistry earth science and physics It guides students in acquiring such skills as finding main ideas making inferences summarizing information and recognizing cause and effect relationships It also teaches students how to interpret scientific data and find the meaning of scientific words from context

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