

# Limiting Reactant, Theoretical Yield, Actual Yield, & Percent Yield

$22.99 + 35.45 = 58.44$



- If we have 53.2 g of Na and 65.8 g of  $\text{Cl}_2$ , what are the limiting reactant and theoretical yield?

$$53.2\text{ g Na} \times \frac{1\text{ mol Na}}{22.99\text{ g Na}} \times \frac{2\text{ mol NaCl}}{2\text{ mol Na}} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}} = 135.23$$

$$65.8\text{ g Cl}_2 \times \frac{1\text{ mol Cl}_2}{70.9\text{ g Cl}_2} \times \frac{2\text{ mol NaCl}}{1\text{ mol Cl}_2} \times \frac{58.44\text{ g NaCl}}{1\text{ mol NaCl}}$$

# 12 3 Section Review Limiting Reagent And Percent Yield

**Alexander Stone Macnow**



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