

KEY

1. The combustion of 1.00 mol of sucrose,  $C_{12}H_{22}O_{11}$ , evolves  $5.65 \times 10^3$  kJ of heat. A bomb calorimeter has a calibrated heat capacity of  $1.23 \text{ kJ/}^\circ\text{C}$ . How many grams of sucrose should be burned to raise the temperature of the calorimeter and its contents from  $23.0^\circ\text{C}$  to  $80.0^\circ\text{C}$ ?  $\Delta T = 80.0^\circ - 23.0^\circ = 57.0^\circ\text{C}$

$$q = C_{cal} \cdot \Delta T = 1.23 \frac{\text{kJ}}{^\circ\text{C}} \cdot 57.0^\circ\text{C} = 70.11 \text{ kJ needed.}$$

$$1 \text{ mole sucrose} = 342.3 \text{ g} ; \frac{342.3 \text{ g}}{5650 \text{ kJ}} = \frac{x \text{ g}}{70.11 \text{ kJ}} ; x = \boxed{4.25 \text{ g sucrose}}$$

\* don't have to do this.

The heat of reaction for burning 1 mole of a certain compound X is known to be  $-477.7 \text{ kJ}$ . The calorimeter constant of the bomb being used is  $2.5 \times 10^3 \text{ J/}^\circ\text{C}$  and the initial temperature of the water is  $23.2^\circ\text{C}$ .

- (A) If  $22.54 \text{ g}$  of compound X (MM =  $46.0$ ) is burned in the bomb calorimeter containing  $2000 \text{ mL}$  of water (S.H. =  $4.184 \text{ J/g}^\circ\text{C}$ ), what will be the final temperature?

$$\boxed{T_f = 44.7^\circ\text{C}}$$

- (B) How much water can be warmed from  $23.2^\circ\text{C}$  to  $56.5^\circ\text{C}$  when  $52.3 \text{ g}$  of the compound is burned in the bomb?

$$\boxed{3300 \text{ g H}_2\text{O}}$$

3. A  $500.0\text{-g}$  piece of metal at  $60.0^\circ\text{C}$  is placed in  $200.0 \text{ g}$  of water at  $22.0^\circ\text{C}$  contained in a coffee cup calorimeter. The metal and water come to the same temperature at  $32.5^\circ\text{C}$ . The specific heat of water is  $4.184 \text{ J/g}^\circ\text{C}$ .  $q_{\text{lost by metal}} = q_{\text{gained by water}}$   $\Delta T_{\text{water}} = 32.5^\circ - 22.0^\circ = 10.5^\circ\text{C}$

- (A) How much heat energy did the metal give up to the water?

$$q_{\text{water}} = (200.0 \text{ g})(4.184 \text{ J/g}^\circ\text{C})(10.5^\circ\text{C}) = \boxed{8786 \text{ J released by metal}}$$

# 12 4 Review And Reinforcement Calorimetry Answers

**Richard B. Kemp**



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