

Chapter 12 Review

Stoichiometry

Important Vocabulary

- Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry
- Mole ratio- a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles
- Limiting reagent- The reactant that determines the amount of product that can be formed by a reaction
- Excess reagent- The reaction that is not completely used up in a reaction
- Theoretical yield- Maximum amount of product that can be formed from given amounts of reactants
- Actual yield- the amount of product that actually forms when the reaction is carried out in a laboratory
- Percent yield- the ration of the actual yield to the theoretical yield in percentage form

Important information

Six steps on how to do stoichiometry

1. Balance the equation
2. Convert to mole
3. Find the ratios
4. Find the limiting reagent
5. Evaluate ratios to find the limiting reagent
6. Convert back to mass

When converting to liters, you multiply by 22.4, but while converting to grams, you multiply by the molar mass

Important Equations

1. Mole conversion= $\frac{\text{Mass of Object}}{\text{Molar Mass}}$
2. Percent Yield= $\frac{\text{Actual Yield}}{\text{Theoretical Yield}}$

Review Problems

Multiple Choice

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Anthony S. Fauci



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