

Mathematically

$$\frac{p^0 - p^s}{p^0} = x_2$$

Here  $p^0$  = Pressure of solvent in pure stage.

$p^s$  = Pressure of solution

$x_2$  = Mole fraction of solute.

### DERIVATION OF THE LAW

Pressure of solvent in solution

= Mole fraction of solvent  $\times$  pressure of solvent in pure stage

**If solute is non-volatile**

Total pressure of solution = Pressure due to solvent

= Mole fraction of solvent  $\times$  pressure of solvent in pure stage

So  $p_{\text{solvent}} = x_1 \cdot p_{\text{solution}}$  Here  $x_1$  is mole fraction of solvent.

$$\frac{p_{\text{solvent}}}{p_{\text{solution}}} = x_1$$

Subtract each term from 1 on both side.

$$1 - \frac{p_{\text{solvent}}}{p_{\text{solution}}} = 1 - x_1$$

$$\frac{p_{\text{solution}} - p_{\text{solvent}}}{p_{\text{solution}}} = x_2$$

### 2.11 IDEAL & NON-IDEAL SOLUTION

| Sr. | Ideal solution                                                                                     | Non-Ideal solution                                                                                     |
|-----|----------------------------------------------------------------------------------------------------|--------------------------------------------------------------------------------------------------------|
| 1.  | A solution which obey Roult's law is called ideal solution.                                        | A solution which do not obey Roult's law is called non-ideal solution.                                 |
| 2.  | Interaction between components of mixture i.e.-between A&B is same as found in A-A & B-B molecule. | Interaction between components of mixture i.e.-between A&B is not same as found in A-A & B-B molecule. |

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