Limiting Reagents and Percentage Yield Worksheet

1. Consider the reaction

 $I_2O_3(g) + 5 CO(g) ----> 5 CO_2(g) + I_2(g)$

a) 80.0 grams of iodine(V) oxide, I₂O₅, reacts with 28.0 grams of carbon monoxide, CO.

Determine the mass of iodine I₂, which could be produced?

SO g LOs	1 mol I ₂ O ₈	1 mol Iz		2028
1	333.8 g I ₂ O ₅	1 mol I ₂ O ₈		
	1 sa sassa 11		253.8 g Iz	
28 g CO	1 mol CO	1 mol Iz	200.08.2	= 50.8 g I ₃
1	28 g CO	5 mol CO	1 mol I,	

- b) If, in the above situation, only 0.160 moles, of iodine, I₂ was produced.
 - i) what mass of iodine was produced?

0.160 mel l ₂	253.8 g. la	= 40 6 o to
	I mol Is	-40.0 E 13

ii) what percentage yield of iodine was produced.

Zinc and sulphur react to form zinc sulphide according to the equation.

$$Zn + S \rightarrow ZnS$$

If 25.0 g of zinc and 30.0 g of sulphur are mixed,

a) Which chemical is the limiting reactant?

= 382 mol ZnS	1 mol ZnS	1 mel Zn	25.0 g Zm
	l mel Zn	65.41 g Zn	1
= 934mol ZnS	1 mol ZnS	1 mol S	30.0 g S
	1 mol S	32.1g S	

b) How many grams of ZnS will be formed?

382 mol ZnS	97.51 g ZnS	= 37.2 o ZnS
	1 mel ZnS	- 37.1 g Zato

c) How many grams of the excess reactant will remain after the reaction is over?

.934mol ZnS -.382 mol ZnS = .552mol ZnS

552 mol ZnS	1 mol S	32.1g S	- 177.0
	1 mol ZnS	1 mel S	= 17.7g a

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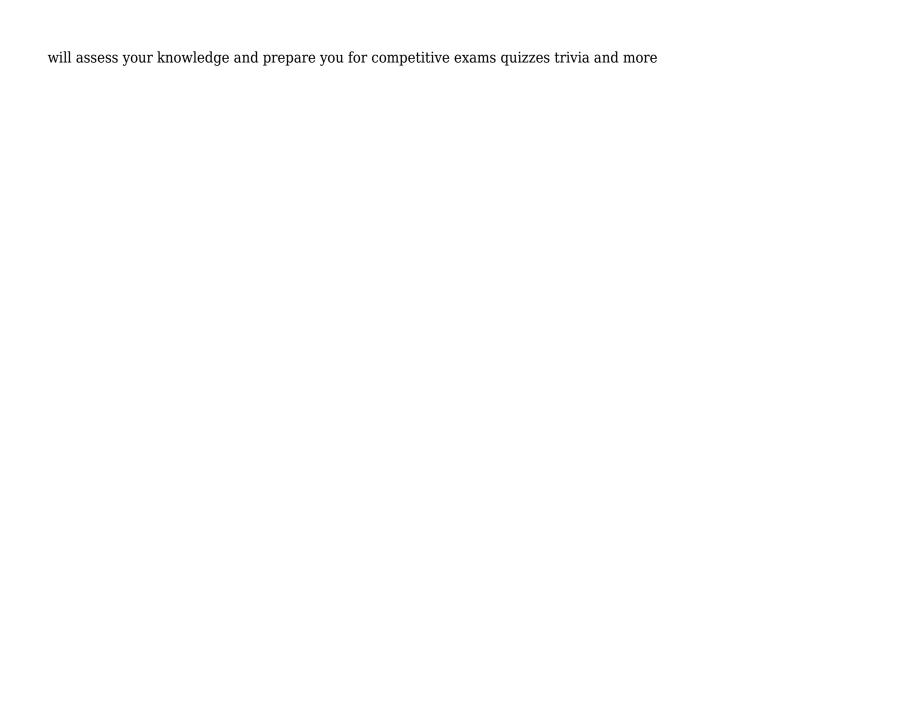
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