

Solutions

Less quantity ← Solute + Solvent → Large quantity
Other than solvent ← (B) (A) → Same physical state as that of sol

$$1. \frac{W}{W} \% = \frac{W_B}{W_{sol}} \times 100$$

$$2. \frac{W}{V} \% = \frac{W_B}{V_{sol}} \times 100$$

$$3. \frac{V}{V} \% = \frac{V_B}{V_{sol}} \times 100$$

$$4. ppm = \frac{W_B}{W_{sol}} \times 10^6$$

$$5. X_A = \frac{n_A}{n_A + n_B}, X_B = \frac{n_B}{n_A + n_B}$$

$$X_A + X_B = 1 \text{ (Always)}$$

$$6. \text{Molarity, } M = \frac{n_B}{V_{sol}(L)}$$

$$7. \text{Molality, } m = \frac{n_B}{W_A(kg)}$$

$$8. \text{Normality, } N = \frac{\text{No. of gm eq. of B}}{V_{sol}(L)}$$

$$= \frac{n_B \times n.f.B}{V_{sol}(L)} \left[\begin{array}{l} \text{No. of gm eq.} \\ = \frac{W_A}{E.W} = \frac{W}{M.W} = n \times n.f. \end{array} \right]$$

9. N-factor / z (valency factor)

- Acid - Ionisable H^+ ion
- Base - OH^- ions furnished
- Salts - Positive or Negative charge
- Ions - Charge anion or cation

Relation b/w concn terms:

$$N = M \times n.f$$

$$m = \frac{X_B \times 1000}{X_A \times M.W.A} = \frac{W_B \times 1000}{W_A \times M.W.B}$$

$$M = 10 \times \frac{W}{W} \% \times d \quad M = \frac{m \times d}{1 + \frac{M.W.B \times m}{1000}}$$

$$11. \text{Dilution Equation: } M_1 V_1 = M_2 V_2$$

$$12. \text{Mixing law: } M_f = \frac{M_1 V_1 + M_2 V_2 + \dots}{V_1 + V_2 + \dots}$$

$$13. \text{For same nature of solution, the resulting Normality} = \frac{N_1 V_1 + N_2 V_2 + \dots}{\text{Total Volume}}$$

$$14. \text{For different nature of solution, the resulting normality} = \frac{|N_1 V_1 - N_2 V_2|}{\text{Total Volume}}$$

$$15. \text{Law of equivalence: } aA + bB \rightleftharpoons cC + dD$$

gm equivalents of any reactant reacted = gm equivalents of any product form

16. Vapour pressure :- Pressure exerted by vapour on liq. surface at equilibrium. $H_2O(l) \rightleftharpoons H_2O(g)$

- * Factors affecting V.P.
- (a) $V.P \propto$ Volatility
 - (b) $V.P \propto$ Temp.
 - (c) $B.P \propto \frac{1}{V.P}$ Note. $V.P$ of Non-volatile = 0

17. Raoult's law for volatile solute

$$P_A \propto X_A, P_B \propto X_B$$

$$P_A = P_A^\circ X_A, P_B = P_B^\circ X_B$$

$P_A^\circ = V.P$ in pure form
 $P_B^\circ = V.P$ in pure form

$$P_s = P_A^\circ X_A + P_B^\circ X_B \quad \text{or} \quad P_s = (P_B^\circ - P_A^\circ) X_B + P_A^\circ$$

18. Raoult's law for non-volatile solute

$$P_s = P_A^\circ X_A \text{ (as } P_B^\circ = 0)$$

$$19. \text{Mole fraction in Vapour phase } (Y_A + Y_B = 1)$$

$$Y_A = \frac{P_A}{P_s} = \frac{P_A^\circ X_A}{P_A^\circ X_A + P_B^\circ X_B}, Y_B = \frac{P_B}{P_s} = \frac{P_B^\circ X_B}{P_A^\circ X_A + P_B^\circ X_B}$$

20. Ideal Solution :- Obey Raoult's law (Generally $\rightarrow$ F.O.A b/w B-B or A-A similar to A-B. dilute soln)

- $\rightarrow$ Properties: (i) $P_s = P_A^\circ X_A + P_B^\circ X_B$ (ii) $\Delta S_{mix} = +ve$ (iii) $\Delta V_{mix} = 0$ (iv) $\Delta H_{mix} = 0$ (v) $\Delta G_{mix} = -ve$
- Examples: - Generally homologous series
(a) $CH_3CH_2Br + CH_3CH_2CH_2Br$ (b) Benzene + Toluene
(c) n-pentane + n-hexane

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Thomas Griffiths



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