

CHEMISTRY 9TH - Short / Detailed Question Answers

→ THE PHYSICAL STATES OF MATTER

CHAPTER# 05

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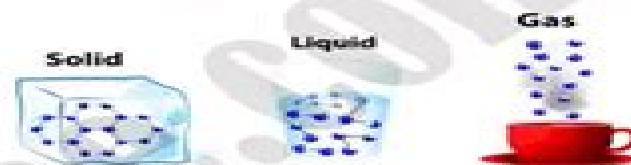
Q.1: Define matter and its states.**Ans: Matter:**

As we know that matter is the physical material of the universe. It is defined as anything which has mass and occupies space.

States of Matter:

These states differ in some observable properties.

- (i) **Gas:** A gas has no fixed volume and shape; it can easily be compressed and expanded.
- (ii) **Liquid:** A liquid state has no fixed shape but has fixed volume it cannot be compressed easily.
- (iii) **Solid:** A solid has a definite shape and volume, it cannot be compressed easily. In addition to the above three states, there are two more states of matter named plasma state and Bose-Einstein condensate. The different physical states of matter are due to the arrangement of molecules and intermolecular forces.

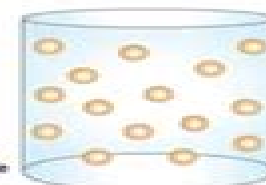
**Q.2: Describe the characteristics of the gaseous state.****Ans: Characteristics of Gaseous State:**

The gaseous state molecules are lying away from one another, this assumption was proposed by Boltzmann, Maxwell, Kelvin. They explain the behavior of gases according to their kinetic molecular theory. Gaseous state shows the following characteristics.

- (i) The molecules in gases are widely separated from each other.
- (ii) The molecules have negligible volume.
- (iii) The gas molecules are in constant random motion.
- (iv) The gas molecules move in a straight line until they collide with each other or the wall of the container.
- (v) On collision molecules do not lose energy because they are elastic in nature.
- (vi) Pressure produced when molecules collide with the wall of the container.
- (vii) There are attractive and repulsive forces between molecules.

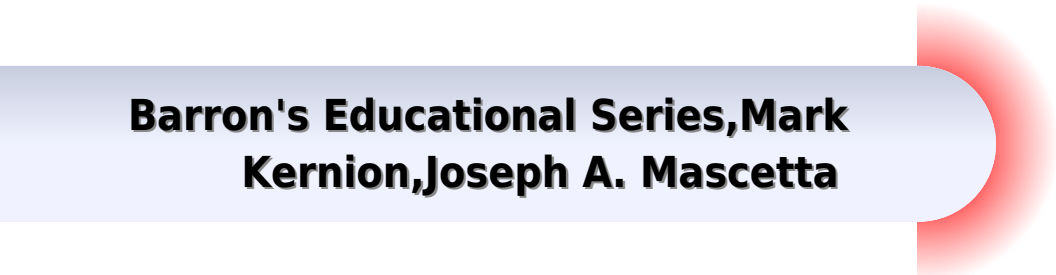
GAS

- not rigid
- no fixed shape
- no fixed volume



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