

13.1**THE NATURE OF GASES****Section Review****Objectives**

- Describe the assumptions of the kinetic theory as it applies to gases
- Interpret gas pressure in terms of kinetic theory
- Define the relationship between Kelvin temperature and average kinetic energy

Vocabulary

- | | | |
|------------------|------------------------|-----------------------------|
| • kinetic energy | • vacuum | • pascal (Pa) |
| • kinetic theory | • atmospheric pressure | • standard atmosphere (atm) |
| • gas pressure | • barometer | |

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

- The kinetic theory describes the 1 of particles in matter and the forces of attraction between them. The theory assumes that the volume occupied by a gas is mostly 2, that the particles of gas are relatively 3, move 4 of each other, and are in constant 5 motion. The 6 between particles are perfectly elastic so that the total 7 remains constant. Gas pressure results from the simultaneous collisions of billions of particles with an object. Barometers are used to measure 8 pressure. Standard conditions are defined as a temperature of 9 and a pressure of 10.
1. _____
 2. _____
 3. _____
 4. _____
 5. _____
 6. _____
 7. _____
 8. _____
 9. _____
 10. _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 11. Atmospheric pressure is 760 mm Hg.
- _____ 12. The SI unit of pressure is the pascal.

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