

CHARLES LAW WORKSHEET #2

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

NAME KRY

1. A gas measures 150 mL at 1.0 atmospheres and 20°C.
Calculate its volume in mL at 0°C and 1.0 atmospheres.

$$V_1 = 150 \text{ mL}$$

$$T_1 = 20^\circ\text{C} + 273 = 293 \text{ K}$$

$$T_2 = 0^\circ\text{C} + 273 = 273 \text{ K}$$

$$\frac{150 \text{ mL}}{293 \text{ K}} = \frac{V_2}{273 \text{ K}}$$

$$\frac{(150 \text{ mL})(273 \text{ K})}{(293 \text{ K})} = \frac{(293 \text{ K})V_2}{(293 \text{ K})}$$

$$V_2 = 140 \text{ mL}$$

2. A gas occupies a volume of 4.50 liters at 27°C. At what temperature in °C would the volume be 6.0 liters, the pressure remaining constant?

$$V_1 = 4.50 \text{ L} \quad T_1 = 27^\circ\text{C} + 273 = 300 \text{ K}$$

$$V_2 = 6.0 \text{ L} \quad T_2 = ?$$

$$\frac{4.50 \text{ L}}{300 \text{ K}} = \frac{6.0 \text{ L}}{T_2}$$

$$\frac{(4.50 \text{ L})T_2}{4.50 \text{ L}} = \frac{(300 \text{ K})(6.0 \text{ L})}{(4.50 \text{ L})}$$

$$T_2 = 400 \text{ K} - 273 \text{ K}$$

$$= 127^\circ\text{C}$$

3. A sample of gas occupies 220 mL at 10°C and 750 torr.
What volume in mL will the gas have at 20°C and 750 torr?

$$V_1 = 220 \text{ mL} \quad T_1 = 10^\circ\text{C} + 273 = 283 \text{ K}$$

$$V_2 = ?$$

$$T_2 = 20^\circ\text{C} + 273 = 293 \text{ K}$$

$$\frac{220 \text{ mL}}{283 \text{ K}} = \frac{V_2}{293 \text{ K}}$$

$$\frac{(220 \text{ mL})(293 \text{ K})}{283 \text{ K}} = \frac{(283 \text{ K})V_2}{283 \text{ K}}$$

$$V_2 = 228 \text{ mL}$$

4. A gas occupies a volume of 90.0 mL at 27°C and 740 torr.
What volume in mL will the gas have at 5°C and 740 torr?

$$V_1 = 90.0 \text{ mL} \quad T_1 = 27^\circ\text{C} + 273 = 300 \text{ K}$$

$$T_2 = 5^\circ\text{C} + 273 = 278 \text{ K}$$

$$V_2 = 83.4 \text{ mL}$$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \Rightarrow \frac{90.0 \text{ mL}}{300 \text{ K}} = \frac{V_2}{278 \text{ K}}$$

$$\frac{(278 \text{ K})(90.0 \text{ mL})}{(300 \text{ K})} = \frac{V_2(300 \text{ K})}{300 \text{ K}}$$

5. A gas occupies a volume of 120 mL at 27°C and 630 torr.
At what temperature in °C would the volume be 80.0 mL at 630 torr?

$$V_1 = 120 \text{ mL} \quad T_1 = 27^\circ\text{C} + 273 = 300 \text{ K}$$

$$V_2 = 80.0 \text{ mL}$$

$$\frac{(120 \text{ mL})T_2}{120 \text{ mL}} = \frac{(80.0 \text{ mL})(300 \text{ K})}{(120 \text{ mL})}$$

$$\frac{120 \text{ mL}}{300 \text{ K}} = \frac{80.0 \text{ mL}}{T_2}$$

$$T_2 = 200 \text{ K} - 273 \text{ K} = -73^\circ\text{C}$$

Answer Key For Boyles Law

Zhe Xu

Answer Key For Boyles Law:

Answer Key For Boyles Law Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has become more evident than ever. They have the capacity to inspire, provoke, and ignite change. Such could be the essence of the book **Answer Key For Boyles Law**, a literary masterpiece that delves deep in to the significance of words and their impact on our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book is key themes, examine its writing style, and analyze its overall effect on readers.

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