

Unit 5 Review: Bonding and Metals

Determine which type of bond would occur between the atoms: ionic or covalent.

Atom 1	Atom 2	Type of Bond
Aluminum	Phosphorus	ionic
Sulfur	Oxygen	covalent
Lithium	Nitrogen	ionic
Carbon	Sulfur	covalent

Write the formulas for the following compounds and then determine if the bonds are ionic or covalent. Draw the dot diagram for the compound. If it is covalent, indicate its shape.

Magnesium Oxide
 Type: ionic
 Chemical Formula: MgO
 Dot Diagram:
 $Mg \cdot \cdot \cdot O \cdot \cdot \cdot \rightarrow [Mg]^{2+} [O]^{2-}$
 Shape (if covalent):

Carbon Tetrachloride
 Type: covalent
 Chemical Formula: CCl_4
 Dot Diagram:
 $C \cdot \cdot \cdot 4 \cdot \cdot \cdot Cl \cdot \cdot \cdot \rightarrow C \cdot \cdot \cdot Cl \cdot \cdot \cdot Cl \cdot \cdot \cdot Cl \cdot \cdot \cdot Cl \cdot \cdot \cdot$
 Shape (if covalent): tetrahedral

Methane
 Type: covalent
 Chemical Formula: CH_4
 Dot Diagram:
 $C \cdot \cdot \cdot 4 \cdot \cdot \cdot H \cdot \cdot \cdot \rightarrow H \cdot \cdot \cdot C \cdot \cdot \cdot H \cdot \cdot \cdot H \cdot \cdot \cdot H \cdot \cdot \cdot$
 Shape (if covalent): tetrahedral

Carbon Monoxide
 Type: covalent
 Chemical Formula: CO
 Dot Diagram:
 $C \cdot \cdot \cdot 4 \cdot \cdot \cdot O \cdot \cdot \cdot \rightarrow :C \cdot \cdot \cdot O \cdot \cdot \cdot$
 Shape (if covalent): Linear

Barium Bromide
 Type: ionic
 Chemical Formula: $BaBr_2$
 Dot Diagram:
 $Ba \cdot \cdot \cdot 2 \cdot \cdot \cdot Br \cdot \cdot \cdot \rightarrow [Ba]^{2+} [Br]^{1-} [Br]^{1-}$
 Shape (if covalent):

Sodium Oxide
 Type: ionic
 Chemical Formula: Na_2O
 Dot Diagram:
 $Na \cdot \cdot \cdot O \cdot \cdot \cdot Na \cdot \cdot \cdot \rightarrow [Na]^{+} [O]^{2-} [Na]^{+}$
 Shape (if covalent):

Bromic Acid
 Type: covalent
 Chemical Formula: $HBrO_3$
 Dot Diagram:
 $H \cdot \cdot \cdot 1 \cdot \cdot \cdot Br \cdot \cdot \cdot 3 \cdot \cdot \cdot O \cdot \cdot \cdot \rightarrow H \cdot \cdot \cdot Br \cdot \cdot \cdot O \cdot \cdot \cdot O \cdot \cdot \cdot O \cdot \cdot \cdot$
 Shape (if covalent): pyramidal

Nitrate Ion
 Type: covalent
 Chemical Formula: NO_3^{-}
 Dot Diagram:
 $N \cdot \cdot \cdot 5 \cdot \cdot \cdot O \cdot \cdot \cdot 3 \cdot \cdot \cdot \rightarrow [O \cdot \cdot \cdot N \cdot \cdot \cdot O]^{-}$
 Shape (if covalent): trigonal planar

Aluminum Sulfide
 Type: ionic
 Chemical Formula: Al_2S_3
 Dot Diagram:
 $Al \cdot \cdot \cdot 3 \cdot \cdot \cdot S \cdot \cdot \cdot \rightarrow [Al]^{3+} [S]^{2-} [S]^{2-} [S]^{2-}$
 Shape (if covalent):

Hydrogen Sulfide
 Type: covalent
 Chemical Formula: H_2S
 Dot Diagram:
 $H \cdot \cdot \cdot 1 \cdot \cdot \cdot S \cdot \cdot \cdot 2 \cdot \cdot \cdot H \cdot \cdot \cdot \rightarrow H \cdot \cdot \cdot S \cdot \cdot \cdot H \cdot \cdot \cdot$
 Shape (if covalent): bent

Use the elements in the following list to answer the questions below: Aluminum, Calcium, Copper, Iron, Lead, Magnesium, Mercury, Nickel, Potassium, Silver, Sodium, Tin, Titanium, Zinc

- Give one of the metals from the above list that is in its ionic form in our good for the body.
Mercury
- Give one of the metals from the list above that can be found free in nature.
Silver, Mercury, Copper, Iron
- Give one of the metals from the list above that cannot be found free in nature.
Sodium, Potassium, Magnesium, Calcium, Aluminum, Lead, Nickel, Tin, Titanium
- Name one of the strongest transition metals from the list above.
Titanium, Nickel, Aluminum
- Give one metal from the list above which is found in nature as a chloride.
Sodium, Potassium, Magnesium, Calcium
- Which metal from the list above gets its name from the Latin word maximum?
Tin
- Which metal from the list above gets its name from the Latin word hydrargyrum?
Mercury

Answer the following questions using one of the transition metals that we discussed: Aluminum, Copper, Iron, Lead, Mercury, Nickel, Silver, Tin, Titanium, Zinc

Property	Transition Metal or Metals
Good Conductors (2)	$Ag, Cu, Al, Fe, or Hg$
High Density (1)	$Hg, Pb, or Ag$
Low Density (1)	$Al or Ti$
Soft (2)	$Sn, Pb, or Hg$
Magnetic (1)	$Fe or Ni$
Good Luster (2)	$Ag, Al, Ni, or Hg$

Uses in Pure Form	Transition Metal
Fuel for Space Shuttle	Al
Car Batteries	Pb
Extreme Cold Iron Masks	Hg
Bike Frames	Ti
Wiring	Cu
Inside of Printer	Zn

Group 1 and Group 2 elements have many similarities and differences. In the table below, give a few words about each of the concepts listed.

Property or Idea	Group 1 Elements	Group 2 Elements
Reactivity with water	vigorous	mild
Density	low density	more dense than Group 1
Occurrence in Nature	never	never
Hardness	soft	harder than Group 1

Match the location with the metal that we need to get from this area.

Location	Metal or Metals
Bismuthville Plant, Utah	Ni
Exton, England	Mg
Dead Sea, Palestine	$Na (K, Mg, Ca)$
Morog, Peru	Ni
Salt Lake City, Utah	$Na (K, Mg, Ca)$

Fill in the chart below about the four types of bonds we learned about. Depends on electronegativity

Bond Type	Particles	Picture	Force	Melting Point	Conductivity	Other Notes	Examples
Molecular (Covalent)	nonmetal	$H \cdot \cdot \cdot H \cdot \cdot \cdot$	weak	low	no	usually gases	carbon dioxide
Ionic	metal ion nonmetal ion	$Na^+ \cdot \cdot \cdot Cl^-$	strong	very high	no (if ionic)	usually solids	sodium chloride
Network Covalent	nonmetal pattern	$C \cdot \cdot \cdot C \cdot \cdot \cdot C \cdot \cdot \cdot C \cdot \cdot \cdot$	strong	very high	no	highly hardest	diamond
Metallic	metal all	$Na \cdot \cdot \cdot Na \cdot \cdot \cdot Na \cdot \cdot \cdot$	strong	varies	yes	ductile malleable	copper

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