CHEM 1A

 Hydrogen has a red emission line at 656.3 nm, what is the energy and frequency of a photon of this light? Note: Planck's constant = 6.626 x 10<sup>-34</sup> J-s, speed of light = 2.998 x 10<sup>8</sup> m/s

$$E_{photon} = hv = \frac{hc}{\lambda}$$

$$E_{photon} = \frac{(6.626 \times 10^{-34} \text{ J·s}) (2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left(\frac{10^{-9} \text{ m}}{1 \text{ nm}}\right)}$$

$$V = \frac{(2.998 \times 10^8 \text{ m/s})}{656.3 \text{ nm} \left(\frac{10^{-9} \text{ m}}{1 \text{ nm}}\right)}$$

$$E_{photon} = \frac{3.02677861 \times 10^{-19} \text{ J}}{4 \text{ m/s}}$$

$$V = \frac{4.568032912 \times 10^{14} \text{ s}^{-1}}{4 \text{ m/s}}$$

Answer (E): 3.027 x 10<sup>-19</sup> J

Answer (v): 4.568 x 10<sup>14</sup> s<sup>-1</sup> or Hz

2. What is the de Broglie wavelength for an electron moving at 2.42 x 10° m/s? Note: m<sub>e</sub> = 9.109 x 10<sup>-31</sup> kg

$$\lambda = \left[ \frac{b}{mei} \right]$$

$$\lambda = \left[ \frac{6.626 \times 10^{-34} \text{ J} \cdot \text{s}}{(9.109 \times 10^{-31} \text{ ke}) (2.42 \times 10^6 \text{ m/s})} \right] = 3.00583657 \times 10^{-10} \text{ m}$$

Answer: 3.01 x 10<sup>-10</sup> m — many orders of magnitude larger than the diameter of an electron

 Using the Bohr equation, calculate is the change in energy when an electron transitions from n = 3 to n= 2 in a hydrogen atom.

$$\Delta E = -2.18 \times 10^{-18} J \left( \frac{Z^2}{n_{\text{final}}^2} - \frac{Z^2}{n_{\text{binal}}^2} \right)$$

$$\Delta E = -2.18 \times 10^{-18} J \left( \frac{1^2}{2^2} - \frac{1^2}{3^2} \right)$$

$$\Delta E = -3.02777778 \times 10^{-19} J$$

Answer: - 3.03 x 10<sup>-19</sup> J

# **Answer Key Study Guide Quantum Theory**

**Karin Nielsen-Saines** 

**Answer Key Study Guide Quantum Theory:** 

This book delves into Answer Key Study Guide Quantum Theory. Answer Key Study Guide Quantum Theory is an essential topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Answer Key Study Guide Quantum Theory, encompassing both the fundamentals and more intricate discussions.

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web jan 1 1997 the lanthanides comprise the largest naturally occurring group in the periodic table there are more than 100 minerals known to contain lanthanides but the only two of commercial importance are monazite and bastnaesite the chapter discusses the production of lanthanide elements by electrolysis of fused salts and by metallothermic reduction 7 1 lanthanoids chemistry libretexts - Nov 30 2022

web the fifteen elements shown in table pageindex 1 from lanthanum la 4f 0 to lutetium lu 4f 14 are lanthanoids ln may be used as a general symbol for the lanthanoid elements although lanthanoids scandium sc and yttrium y are sometimes called rare earth elements they are relatively abundant in the earth s crust

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web feb 4 2021 the lanthanides are located on the top row of the two rows of elements found below the main body of the periodic table placing the lanthanides and the actinides the row below the lanthanides below the other elements is just a matter of aesthetics the lanthanides belong to period 6 like cesium barium and hafnium